

Continue

Bard early college high school

In physics, the gas constant is defined as the product of pressure and volume. Rated with R and expressed as negroids and the value of the gas constant can be expressed in several units. The value of at at my which is at the standard atmospheric pressure is a R-3144598 1.7. *Incl* 1.3414598 2.7. *Incl* 1.3414588 2.7. *Incl* 1.3414588 2.7. *Incl* 1.3414588 2.7. *Incl* 1.34

K) n = 0.391 mol Example #2: A gas sample at 28.0 °C has a volume of 6.20 L and exerts a pressure of 720.0 mmHg. How many gas moles are in the sample? Solution: 1) PV = nRT: (720.0 mmHg/1.00 atm) (6.20 L) = (n) (0.08206 L atm / mol K) (301 K) Note conversion from mmHg to atm and from Celsius to Kelvin. 2. Calculate: n = 0,238 mol Example #3: Calculate the approximate volume of a gas sample of 0,400 mol at 11,0 °C and a pressure of 2,43 atm. Solution: 1) PV = nRT: (72.0 mmHg. How many gas moles are in the sample? Solution: 1) PV = nRT: (720.0 mmHg/1.00 atm) (0.08206 L atm / mol K) (284 K) 2) Calculate: (0,400 mol) (0.08206 L atm / mol K) (284 K) V =

-----mmHg and a volume of 6.00 L. Solution: 1) PV = nRT: (780 mmHg) (6.00 L) = (0,300 mol)

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