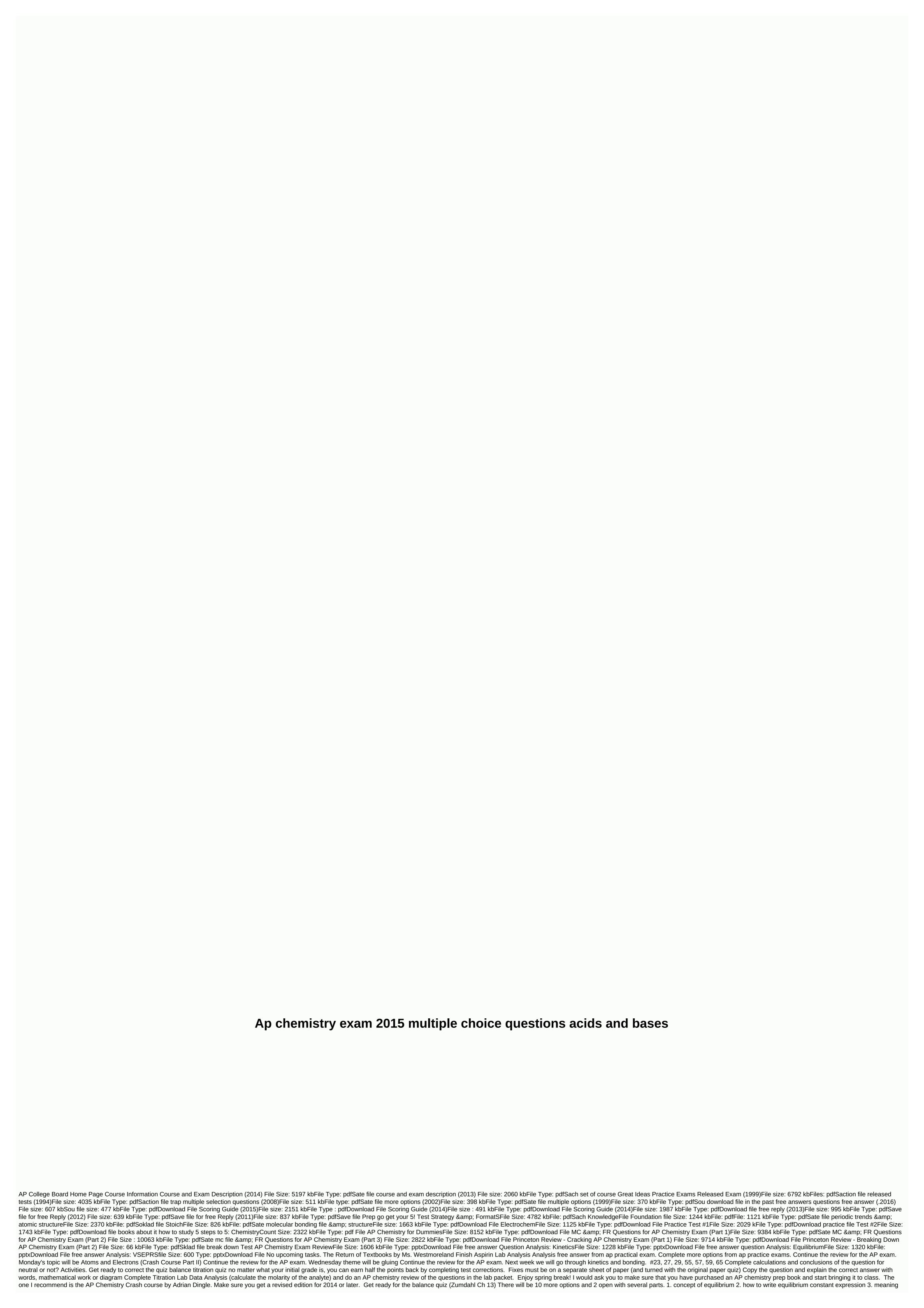
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of value K 4. how to manipulate values K (reverse, change of schimiometry, addition of equations) 5. how to change from Kc to Kp and vice version or restore Lab 1 balancing. Read background, security and procedure 2. Complete Pre-Lab Questions Lab_Restoring Balance.pdf Hellog section with a choice of multiple answers.0010Be forward to answer the questions in the section with a choice of multiple answers, I want to talk and you will be deducted a quarter point for those that you miss.0020Now, you should know that you could actually get a raw score (which is the 5.0033You can answer 35 or 40 only to questions with a choice of multiple options and still end up with 4; so I do not think you have to answer all the started.0067 We'll pass the 1999 AP practice exam; so let's hope you have it in front of you; I have it in front of me here.0071M justice of the started.0067 We'll pass the 1999 AP practice exam; so let's hope you have it in front of you; I have it in front of me here.0071M justice in the started of the st	and welcome back to Educator.com, and welcome back to AP Chemo you, just briefly, about it.0012Such, you will have about 90 minutes number you get correct, minus the number that is missing) about 50; solutely every one.0048So if you are not quite sure about one, go on to go through it; OK.0079The first question gives you the choice of a	stry.0000Today we'll actually start going through the AP exam of to do 75 questions with a choice of multiple answers; you will ge so you could literally skip one third of the questions with a choice chances are your score will still be high enough to keep you in activation energy, free energy, ionization energy, kinetic energy	completely.0004 We're about to start with a set 1 point for each one that you get right, se of answers and still end up with the range of 4-5.0058 He said let's go and truss energy: so these are your
options and we will The questions that follow actually ask you to find out which energy is being talked about.0082 #1: The energy needed to co state gaseous atoms; you come across it with a certain amount of energy and convert it to the gaseous phase of ion.0118You knock down one e electron).0129This is it; #1 is C.0141OK, #2: Energy change that occurs when an ionic solid is converted into a widely separated gaseous ion; O compound together.0160When you have, say, sodium and chlorine, they merge; separately from chlorine; each of them steal the electron from s merge, they are actually more energy-stable - remember, we said that all systems are looking for the lowest energy.0180Ionic compounds, due to	ectron and it is written as follows (this is the first ionizing energy - the C, now it's one of the things we haven't formally discussed in our particular, and then you have chloride ion, which is negatively charged, a positive and negative charge, release of energy; so it is grid energy.	second ionizing energy and so on, it will only be this thing when cular course: it's called grid energy.0143A union energy is essented sodium, which is positively charged.0167Combine, because point is the amount of energy that is necessary to keep it together.	you knock down another Itially the energy that keeps an ionic Dositive and negative attract; when they D192In order to divide it - to separate it as
ar away as possible - to separate the sodium ion from the chloride ions - which is called grid energy, because these ion compounds consist of a DK, this is an important word here, useful work: this is free energy.0219OK, that's B; and you remember free energy: the reason we called it free options.0243Such I would say a good 75% of the questions with the choice of multiple options are things either you know or do not know and carneed.0257Every calculation you need to do are actually really, really simple numerical calculations - things like .5, .25, 2, 4 - numbers that you can ave to go through the transition state: it is called activation energy.0275The activation energy is A, so 4 is A.0287OK, let's see - let's move on.0	energy - it's the energy that's available for real work, not random ener answer quite quickly.0248There are some calculations; Now of cours n multiply very quickly.0264OK, #4: Energy needed to create a transic	gy, which would be heat.0231OK, so they're quite simple; and a se you can't use the calculator in this part of the exam - all you g ent state in a chemical reaction: so you remember that when we	gain, that's how it will be on more exam et is a periodic table - but that's all you switch from reactants to products, we
which one is which.0299 We have to identify.0314 #5 says it represents an atom that is chemically non-reactive; well, if you look at the options 2p6.0332Death on #5 is D.0338OK, represents an atom in an excited state: so, in an excited state excited means that the atom normal electron been ejected from its earth state energy level (what it may be) to a higher level; so it will have orbitals lower than that, it does not need to be filled atom that has 4 valence electrons: OK, 4 valence electrons – remember that valence electrons are all electrons in the highest primary shell.0381 electronic level 2: 2s, 2p: there are 2 in s orbitals; there are 2 in p; it gives you a total of 4 valence electrons; so #7 is C.0405OK, #8: represents the content of the content	configurations when we make them - they all go to the lowest energy .0355No, the only one that satisfied that one is: you have 1s, which is / here we have let's see: 2s, no; B is 2s2, 2p4 No, I'm sorry, 2s2, he atom of the transitional metal: well, the transient metal is the one w	r; we go from the lowest energy, and we work our way up.0342 E empty; you have 2s with one electron in it; it is excited state ato 2p2; there you go.0393C is the answer here, and the reason is before the orbits d are actually filled - in the periodic table that the	Excited state means that the atom has m.0366 #6 is A.0375 #7: represents an pecause valence shell is the primary middle part with 10 columns.0423This are
ransient metals; so here, looking for something filled d orbitals, and that it will be E.0436 # 8 is E; OK.0445Let see: so now questions 9 to 12 coracid and sodium dihydrogen phosphate; hydrochloric acid and NaCl; NaOH and NH3; and NH3 and acetic acid.0465This are our steam a solu DK.0488 Now, the lowest pH means the most acidic.0494OK, now let's look here: as far as the most acidic concerns, the only thing I see that is fact, they do nothing.0518Neus will not act as a foundation; will not affect the equilibrium concentration of hydrogen ion; so in this case, the answ neutral solution will be E, NH3 and acetic acid.0544All right, you might think it well, let's just a little go through the options.0561It will not nach	ion with a pair of these species; these are our choices.0479A again some strongest acid is HCI.0507A then we look at NaCI; well, sodium cher is C.0529Nadej give me a small box around it: so the lowest pH = tand nh3, because they are both bases; it will not be C, HCI and NaCI,	aid that they are in the ratio of 1:1 mole.0485Sto, #9: we are loo loride, when they actually break down in a solution for sodium ar the most acidic.0539OK, #10: the most almost neutral solution: C because it is acidic; H3PO4 is a weak acid and sodium dihydrog	king for a solution with the lowest pH; nd chloride - they will be spectator ions; in DK, in this particular case, the most almost gen phosphate that you break
up.0568H2PO4, it is also a weak acid, so you are probably looking at an acid solution.0586NH3 and NH4CI: well, this - you might think it's A, but dissociation (or association) balance of weak base - in this case NH3.0609Well, we know (or we should know) that Kb for this is much smaller that atio of 1: 1, so even if there will be some dissociation, Kb is really, really small so it will not be so much.0654So, for all practical purposes, amm somewhere in the order of 10-5 and 10-6).0664When you have a negative logo that you will have pOH.0681No, pOH will not be anywhere near reand gives you a pH close to about 7.0685So, E is the best answer to #10.0707OK, see: #11: Buffer at a pH greater than 8so looking for a buffer you have weak acid and its conjugate base, you get a solution that is acidic.0760It will pH less than 7; we do not know where I think we can ad	n 1.10628Well, Kb for this is equal to the concentration nh4 +, times to onium concentrations and ammonia concentrations will be the same, seutral; but E, where you have ammonia (which is a weak base) and a or, which is basic.0711Lette here (remember what is the buffer?): Buffer	he concentration oh- divided by the concentration nh3.0637But so what you have left is: the concentration of hydroxide is actuall cetic acid (which is weak acid), are in a 1:1 molar ratio, so one are is a weak acid plus its conjugated base, or it is a weak base pl	they said that NH3 and NH4Cl are in a y going to be equal to Kb (which will be actually ends up neutralizing completely us its conjugated acid.0730It is used when
will be greater than 8, so you are looking for a weak base and its conjugated acid.0775 As it turned out, And it satisfyes; so 11 is A: the weak bas acidic.07960ds. his conjugate base Remember how we get a conjugate base - we simply pull out one of the hydrogen ions; it will be H2PO4-, of bonding in different types of solids.0844Key word is fixed here. 08500K, well, let's just pass; so #13: caesium chloride - caesium chloride is an always be so that for ionic compounds: bonding in an ionic solid is actually quite easy – it's just positive and negative bullets stuck together – so this: one of the things that makes metals so good conductors of electricity is precisely because electrons are delocalized.0903So, when I have a	e is ammonia; its conjugate acid is ammonium; chloride is irrelevant.0 and this is your conjugate base.0819B is a combination.0831OK, let's ionic compound; is made of cesite cations and chloride anions; and it o worries.0880 #14: Gold – OK, so gold is metal and consists of a tig	784OK, #12: Buffer with a pH of less than 6: so now it's the one see: let's move on to questions 13 to 16.0833OK, so questions 2 will be A.0854Is a grid of positive and negative ions held togeth htly wrapped grid with delocalized electrons throughout.0892 #1	- it's the buffer - that will be 13 to 16 refer to the following descriptions er by electrostatic forces.0875So, it will 4 will be B; so what we mean when we say
electricity, because electrons can actually move by sample (or across the surface of the sample,0928When you see delocalized electrons that re would be: Strong more covalent bonds, including pi (yes, we have pi bindings here; double beds represent one binding to pi), with weak intermolecular forces.0983A part of the thing: if you look at methane, if you remember 1000 you get nonpolar molecules.0989They are the only bindings, so the best answer would be C.1007OK, so questions 17 and 17 18 refer to the follows degrees Celsius).1025It is necessary fluorine.103317 is E, fluorine.1038He now, reacts with water to form a strong base: lithium, nickel, brong the content of the thing is the delocalized	present metal bonding.0938OK, #09 carbon dioxide - okay, carbon diocular forces.0958Noho, CO2 is a non-linear molecule (right?ductopine Lewis structure for methane, you have something that looks like a wing elements: we have lithium, nickel, bromine, uranium and fluoring	oxide - well, if you remember carbon dioxide Lewis structure for c moments - in fact, it is disturbed: it is non-dollary), so this wou small tripod here, and dipole moments of each individual CH books. 2.1014 #17 says: Is the gas in standard condition at 298 Kelvin (or carbon dioxide is it.0944 Our choice ld be D; this is the best option.096816: We nd are canceled, and when you cancel, 298 Kelvin's room temperature, only about
ithium, sodium, such things can not really be stored in water, because they will react.1053 Let's go ahead and put 18 here; 18 will be lithium, who take lithium metal, and if you add it to the water (which I'll write as HOH), what ends up happening is well, literally, what ends up happening ithium.1085 They form a hydrogen gas (they merge to form a hydrogen gas); leaves 2 hydroxides free, and therefore what you end up with - lithing the properties of the statements below is followed by 5 suggested answers or completions; choose the one that is best in any case, and then poptions, C is the best: the spark plug supplies part of the activation energy for the combustion reaction.1167The departure reactions are highly experienced by the combustion of the statements are highly experienced by the combustion of the statements are highly experienced by the combustion of the statements are highly experienced by the combustion of the statements are highly experienced by the combustion of the statements are highly experienced by the statement and the statement are statements are statements.	ch is A, and just so you know what the reaction actually looks like it is (I'll draw 2 of these and 2 lithium): a hydrogen atom - one of these im + and hydroxide minus - merges; you end up with lithic hydroxide fill in the corresponding oval on the answer sheet.1137OK, so which o	s actually why they call them alkaline metals because they form hydrogen atoms - steals an electron from lithium; another hydro n the solution, which is basic; that's what happens.1114OK, so lof the following best describes the role of spark from the spark pl	hydroxide in the solution.1068As long as gen atom steals another electron from et's see: Starting with #19, each of the ug in the car engine?1159OK, of these
that reaction happens.1185My have very high activation energy; spark gets activation power.1196If I put oxygen and hydrogen in a container and the reaction forward.1212Rection is here C; 19 is C.1218OK, what weight Au is produced when 0.05 moles Au2S3 is completely reduced with expur reaction will be Gold is formed, and then it will be H2S, and then we will have to balance this equation out, so we have ourselves 3, 2, 3; I to produce let me see, so for every 1 mole, what weight Au?1273For each mole from it we produce 2 moles (right? for every 1 mole, we produce which again - you will have a periodic table available to you, so you will have molar mass, and this is a really simple calculation): .1 times 197 is	do nothing, well, they want to make water really badly, but they can't cess H2?1224OK, so let's write a response to this: so we have: Au2Sink it will be 3, 2 and 3 (3 H2, 3 S, 3 S, 2 Au); Yes, this is our balance to 2 moles) of gold.1295.05 moles it produces 0.100 moles of gold; world 19.7 grams.130119.7 grams is the answer; which makes it B.133	get past the activation power.1203When I light it - when I make is reduced completely with a surplus of H2.1235Mase Au so d equation.1247OK, so say that we have 0.0500: 0.0500 moles ell, now we do 0.100 moles (and again, these numbers are very, 3Go good; OK, now that the sodium chloride solution evaporate	only a small spark - that's it, what makes they say that gold is indeed formed; So Au2S is completely reduced, so we want very easy), times 197.0 grams per mole in flames, the color of the flame is
yellow.1339This is just one of those things you should know: when we do flame testing (we didn't actually discuss it; it's part of a qualitative analy you have to sort of quickly test it to make sure what is there.1350 As it turned out, the sodium ion burns yellow when you dip a piece of metal (paight, yellow light.1386July burns yellow, lithium burns red, and potassium burns blue; so just a little bit I know that.1396A again, it's just one of the we're looking for something that goes into a smaller number of gas particles, or maybe from gas to liquid, gas to solid, liquid to solid such thing moles of oxygen gas that are suddenly contained in this lanthanum oxide, so you have gas to go into solids; so yes, our best option is E, for #22.	per clip or something) into the solution and put it in a flame.1374Hor wase things that just comes.1404OK, #22: Of the following reactions that something that more ordered.1420The drop in entropy means some 1445OK, let's see: #23: hot air balloon, shown above, rises; which of the company of	vill be really yellow, because heat is exciting electrons up to high at include the greatest decrease in entropy?1410OK, so let's see thing that becomes more ordered.1441Z of them, it looks like E he following is the best explanation for this observation?1467Ok	er energy levels; fall back. and they emit e: the greatest decrease in entropy So is the best option because you have three K, the best explanation for this observation:
the air in the balloon is surrounded on the outside by normal air; things rise when their density is less than the density of the surrounding medium because the air density inside the balloon is less than the density of ambient air.1511This is it; so 23 we have E.1518OK, #24: The safest and sodium bicarbonate solution.1533It will be D; the reason it is is: you can wash it down with water, how to get rid of as much of it as possible, but it solution or something like that), because you risk going over the top.1554A strong base also damages in the same way as acid; they actually will then wash it again.1573So, D for #24.1578OK, #15 The cooling curve for a pure substance that changes from liquid to solid is given above; solid	nost effective emergency procedure to treat acidic spray skin is to do there is any excess, you want to use bicarbonate, which is a very, ve do the same damage, so you want to use a weak base, such as sodi a and liquid coexist in this - the answer is C.1581 The answer is C b	which of the following immediately?152424 will be: Rinse the affiry weak base to compensate for acid.1540You want to use any um bicarbonate (baking soda).1562It is best: wash; put in a little ecause all points between Q and S Okay, this is a heating cur	fected area with water and then dilute with strong base (such as sodium hydroxide baking soda - it neutralizes everything; ve. so basically what it tells you is:
whenever you see a straight line on the heating curve, it means that it makes a transition.1601By a straight line, I mean a horizontal line where the decreases; suddenly hits Q, jumps to R and then stays horizontally from R to S.1625This means that there is a change in phase; that is - and beconsoled so solid, both liquid and solid coexistence; all liquids are converted into solids that coexist; so surely coexist between R and S.1645Well, this part have time to completely arrange in the grid necessary for the solid.1657You can actually drop the temperature below the normal freezing point; it form.1680 As ice begins to form, the temperature of this ice actually begins to rise back until it reaches a temperature where it is a normal freezing point; it because we would go very slowly, very slowly; you would just see a drop - or from your point of view a drop, horizontal and then another drop.17	ause we do a drop in temperature, that is, the liquid goes into solid.16 from Q-do-R - as it turned out, you can drop the temperature, drop the is called supercoo cooling.1675All, when you drop to the temperature g point.1692This small cooling part - you will often see that in the hea	633A of course we know this, because that is what it asks.1642A temperature, drop the temperature of the liquid; and if you drop of cooling, at some point they begin to change; so between Q atting curves - at least experimentally, you will see it.1704In ideal	A coastline when we transition from liquid to it kind of quickly, the molecules do not and R, you still have that is, ice begins to conditions, you would not see this cooling,
only if when in fact it is completely solid, completely liquid, or completely gaseous - that is, when the temperature varies depending on its thermal coefficients are reduced to the lowest intempocal terms, the coefficient O2 is 1745Nomach right, so let's do this: 26.1754Mme: C10H12O4S, planead and do carbon first; so we glue here 10 carbon.1782I will give hydrogen another, which is 12 hydrogen; so we put 6 there.1789A then Non the right; we have 2 here; we have 4 here.1812I want 28; if we get 28-4, that's 24; How do we get 24? 12 here and what are they actually also of the visible spectrum of the photometer belongs, which of the following?1842Deminating concentration of nitrate solution copper (copper (2	capacity.1729 We used the discussion of heating curves.1739O in the object of the capacity.1729 We used the discussion of heating curves.1739O in the object of the capacity o	s case yes, C would be the best answer.1742OK, let's see: Whe, sulfur will also oxidize.1759It will react with oxygen, so you will n: oxygen - we have 20 oxygens, plus 6 oxygens, plus 2 oxygen OYou go: just a direct problem with balancing - nothing too stran	en the above equation is balanced and all end up with SO2.1777OK, so let's go as, a total of 28.1794 We have 28 oxygens ge.1838OK, #27: let's see, the appropriate
1855A visible light spectrum photometer is used to determine the concentration of a certain species, so it would be 1 only, so 27 - it would be A.1 course.1886Coordination chemistry usually comes at the end; some teachers talk about it; some no.1899 If you see any questions about coording 2.1904Sto, the reason I did not want to discuss it is because again it is one of those things that, running through the practice exam, you can actus same reason why I decided not to discuss organic chemistry.1927Is attacked what comes up in organic chemistry, it will be a maximum of 2 questable the solutely important and essential.1946You should know: transition metals like copper – basically happens: when - if - you dissolve this copper reason.	B71OK, by the way, I should let you know so that we're actually reaction of chemistry, whether it's naming or something like that, it can be ally get a little information you need in terms of coordinating chemistry tions; and again, if you miss them - if you don't know them - it won't he trate in water, copper nitrate dissolves, but what ultimately happens is	Ily talking about it because I didn't discuss coordinating chemistre anywhere from maybe 1 to 3 questions well, 1 question - real.1915 I did not want to spend the whole lesson or two lessons of urt you at all.1934 I didn't want to spend a lot of time on the less that copper water molecules actually organize around a cop	ry in this particular AP chemistry ally, you will not see more than 1 or n it; it was not worth it, which is for the on; I wanted to focus on things that were per ion: maybe 4 of them, maybe 6 of
them, maybe 2 of them.1955Transitional metals are notorious for absorbing visible light, which is why transition metal solutions actually have difference they actually absorb, based on their color; and this tells us concentration.1991This is a bit of a secondary explanation of what is happening here 2005Magnesium is more positively charged than sodium; O2- is more negatively charged than F-; O2 ion is smaller than F- ion.2020March, spears ionic.2033Well, the stronger the bond, the more energy I have to put into it to break it.2048O, as it turned out, here it will be B: 1 and 2: magner positive charge, more electrostatic force, stronger bond, higher melting point; so it will be B.2061OK, 29: let's see, the organic compound respond to the charge in the properties of the propert	2000OK, 27let's see, 28: OK, the melting point of magnesium oxide king of melting point, we are talking about fees in between; So, if I have sium is more positively charged; oxygen is more negatively charged.2 presented above is an example it is keytone.2079Again, it is so it	is higher than the point of sodium fluoride; explanations of this over magnesium oxide and I want to melt it, it means I'm trying to be 052Magnesium is more positively charged than sodium; oxygen t will be E, and just so you know, keytone consists of a carbonyl	observation include which of the following? oreak the magnesium-oxygen bond, which is more negatively charged than fluoride; group (which is this thing) associated with
wo this carbon is connected to 2 carbons.2088There are more carbons, but it is connected to at least two carbons.2107It is called carbonyl grown will continue on - you will discuss organic chemistry, for those of you that are biology of large companies and chemistry disciplines.2122The exidative numbers, mostly; so let's go ahead and calculate some oxidation numbers here.2144Thhous + 4 O2F2 goes to SeF6 + 2 HF + 4 O2 (I I 22170O2F2 it is a very unusual compound; the only atom that is actually electronegative than oxygen is fluorine, so in this case oxygen will not see that the control of the	e is no problem - no worries that.2135OK, #18 Which of the following ope I copied that correctly).2151Letme just do some oxidation number have an oxidation state of -2181 Fluorine will have an oxidation state ide), fluoride is -1; there are 6 of them, so the total fee -6; that is, sele	is true in terms of the reaction represented above?2139This sturs here: hydrogen will be +1; Selenium will be there are 2 of the of -1; and since there are two of them, the total fee will be -2.21 nium is +6.2213It will be -1, +1 and 0 on it; so in this particular c	ff They seem to be asking about nem, so it's +2, so the selenium here is - .92Well, it must be balanced +2 because it ase it looks like it will be D: the oxidation
number of selenium will change from -2 to +6, so our answer is D.2229OK, let's look31: If the temperature of the aqueous sodium chloride solution remains unchanged - no.2264Density - as you heat something, it will expand, so the density will decrease; Molarity of the solution remaind dissolved substances per kilogram of solvent; the mass of water does not change - its volume varies.2285Meal of solution (water) does not change he pen OK.2308Yem you want me to go ahead and draw a propene atom: now again, that's one of those questions so we did not discuss of maybe.2319Largely means that you have three embers there; each carbon will be attached to each other; yes, it's possible, but let me go ahead and you have a single bond all the way here (so sp3 hybridized).2361In this case it will be sp2 and sp3 hybridized; so it looks like the correct ans	is it will be which means that the total volume changes, so that the rige; its volume varies, so C is the answer here.2295A change changes rganic chemistry; it is possible that a molecule like propene will actual and do it.2343Propene will look like this.2351So basic what you have	molarity does not remain.2274Thisality of the solution remains un , molarity changes OK, that's 31.2305Let to #32: let's see, typ ly appear in lewis's problem with the structure you studied, perha is a double bond that will be sp2 hybridized (whenever you see	nchanged - yes: molality is the foam of bes of hybridization exhibited by C atoms in aps, in the classroom, or in problem sets?- a double bond, atoms are hybridized sp2)
contains .1 moles of sodium chloride and .1 mole of calcium chloride.2391What is the minimum number of Moles agno3, which must be added to one source of chloride.2414 Our other source is CaCl2; it will disociate into calcium 2 +, plus 2 Cl2425Sto, for each mole from the fact that it will eave it like a mole for a liter so we end up with 0,1 moles of sodium chloride solution.2439.1 moles from it produces .1 moles of chloride ions; \$0.1 moles per liter equals 0.1 moles of calcium chloride.2476Well, one mole of which produces 2 moles of chloride; so what I ended up with is 0.3 moles AgNO3?2510Nopneal, I need .3 moles Ag, and AgNO3 one mole that 1 mole Ag + plus one mole NO3-; so I need 0.3 moles AgNO3.25	the solution in order to prey all chloride like AgCI?2403Assume that A be 2 moles chloride; so we need to calculate some moles here.24343 to I have 0.1 mol of chloride ion that comes from sodium chloride solumoles chloride ion.2496Um I end up with a total of 0.3 moles chloride	AgCl is insoluble.2411OK, so basically, we have this: NaCl is distance the second seco	tancing itself into Na + + Cl-, so it will be on't have to do That's 1 liter, so I can ad Oh, the same thing; OK, 1 liter times estion is: What is the minimum number of
magnesium chloride chloride: you need to find the total number of moles of chloride ion, if you are trying to take it all from the solution.2550Slast electrolyte cell, which includes the following half-response.2579This: which of the following so we have AIF63-, plus 3 electrons, goes into alum esponse: AIF63- is reduced to cathode - yes, oxidation occurs on an; reduction occurs on the cathode.2609 Equation as written, which is AIF63- reduction because the electrons are on the left side.2639This is the reduction; reduction happens on the cathode, so our answer to 34 is A.26510 of aluminum produced?26591 faraday equals 96,500 coulombs: OK, this is an electrolyte cell; this is where we will pass a specific stream through	s 1:1 with silver and chloride, and silver comes from silver nitrate, which inum metal, plus 6 fluorides ions.2588Sto, in this case it looks like alures 4 electrons goes to aluminum + 6 F-: this is a decrease.2621 The obt., #21 A steady stream of 10 amps passes through the aluminum pr	ch is 1:1.2563That is it; so the answer is C.2572It is #33; OK.25 minum has been reduced from aluminum 3 + to aluminum 0.260 xidation state of fluorine does not change; aluminum is +3; this is oduction cell for 15 minutes; which of the following is the correct	76#34: Questions 34 and 35 refer to the 101Which of the following occurs in s 0; it's a reduction and we know it's a term for calculating the number of grams
seconds in 1 minute to cancel the minutes; We'll have 10 amps, so it'll be 10 coulombs per second, times 1 mole electron is 96,500 coulombs.26 divided by 96,500, and 3 on the denominator; so 35 is C.2737OK, now 36: initial speed data in the table above were obtained for the reaction below. 2773 Well, actually, you know what, let me go over it a little bit.2781I'm going to check out I start with nitrogen oxide, and I notice experimentary the speed solution of 5.0x10-4; So this is when I double the concentration, the speed doubles.2804Noho, that is, it is the first order in NO.2819 Now I go to 1.1 concentration of O2, the speed is 5x10-4; and if I multiply it by 4 and go to 0.4 for the initial concentration, my speed goes to 8.0x10-3.2836	ow; what is the experimental rate law for reaction 2 NO + O2 gas goestent 1 is a .1 molar initial, and experiment 2 is .2 Molar; so that's the inthe second reactant; I go to oxygen, which is in the second column; so the interest of the second column; so the second	s to NO2?2753OK, so we know that our rate law the speed winitial rate, so we use the initial rate method says that for .1 Molarso now, I use experiment 2 and 3 because I will keep no concente speed by 16, which means that it is 4 squared.2857Thfore, it is	Il be equal to k times the concentration of the initial rate is 2.5x10-4; and for .2 tration fixed.2824I warning when I get .1, the second order in O2; so this means
that our answer is B: speed is K times NO times O2 squared.2868OK, #37: Ionizing energy for element X are given in the table above; based on the fourth and fifth: the jump from 2740 to 11 600 is huge.2892So, chances are, the best thing it represents is probably aluminum; and the reason order to begin to take away these two electrons; and 1815 and 2740 represent a pretty significant jump.2940A then, of course, the fourth and fifth Lewis acid if it receives a pair of electrons to form a bond; that's just something you should know.2966A Lewis acid is something that receives a fand liquid phases at the normal melting point?2983Normal melting point - atmospheric pressure is 760 millimeters, 1 atmosphere; so we go to possible that have the largest dipole moment?3017D - this is the one that is ionic; What's the dipole moment Well, HF some peop	why aluminum is due to the following.2916The 1s2, 2s2, 2p1 the f – 1s2 electrons are just too high – it's too much energy, so just based ew electrons to form a bond.2978OK. #39: Above is shown the phase int C: this is where the phase difference between solid and liquid - sol	irst ionizing energy that takes this electron is quite low.2928Nož d on these numbers, aluminum is the best choice.2952 #38: Let' diagram for a clean substance; which point on the diagram corr id on the left side, liquid in the middle, gas on the right.2998O, C	going here, it will be a significant leap in see, a molecule or ion is classified as esponds to the balance between the fixed is the answer.3013#24: We have: Of the
s non-linear; O2 is non-linear; F2 is non-linear; so it's between A and D; D is actually bigger - fluorine is the most electronegative.3039OK, and # ower value, compared to its value in the original balance?3061OK, I am in balance, constant pressure; In this reaction vessel I put o2.3070Wher Keq remains the same.3086Eas position can change; in other words, SO3, SO2 and O2 different concentrations; but Keq does not change.3086Eo3 gas does not go down; the amount of O2 in response does not reduce because you are injected with O2; the amount of SO2 in the reaction out because the reaction will move to the left.3125Sto, 41 is E.3139OK, thank you for joining us here at Educator.com for this particular discu	11: After mentioned above, some pure O2 gas is injected into the equal put in this reaction vessel some O2, the principle of Le Chateliera - 5Abrasive pressure in the reaction vessel let me see what they say vessel - yes.3114Due the reaction will move to the left, SO2 will be conversed.	ilibrium vessel at a constant temperature.3049Po the balance is it shifts to the left.3079O, what it actually does: keq changes for ? the following has a lower value?3104Ne, the total pressure donsumed - the excess you are using; so2 concentration that was	determined which of the following has a reaction?no, at a given temperature, oes not drop; it actually goes up.3110All there before something happened - it will
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