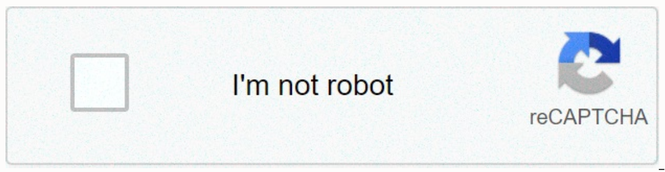

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Determining the chemical formula of a hydrate lab answers

Academia.edu uses cookies to personalize content, customize ads, and improve the user experience. By using our site you agree to collect information through the use of cookies. For more information, check out our privacy policy.× recently did a lab at our school, but some of the questions related to the lab confused me. The data obtained from the lab is here: Determining the chemical formula of a HydratePurposeFind the molecular formula of the hydrate of Copper (II) Sulfate, CuSO4 x H20. PredictionWhen the solution is heated, the hydrate is converted into a water-free ionic compound. Next, the percentage composition is determined by weighing the mass of the white crystals. Materials: 400ml cup with hot plate• Tang• Scoopula• Electronic balance• Glass rod• 3g to 5g hydrated Copper(II) Sulfate• GogglesProcedureThis experiment was done to find the percentage composition of a hydration compound. The mass of the empty beaker, cup + hydrated Copper (II) Sulfate, and mass of the beaker + water-free Copper (II) Sulfate was found by this procedure. The mass of a cup and stirring rod was found. About 4g of hydrated Copper (II) Sulfate was added to the beaker and the weight was recorded. The beaker was placed on a hot plate, heated until the blue crystals lost there blue color. Stirring the compound was essential to ensure that all blue crystals were heated. When the crystals lost their color, the heat source was removed and the beaker was left to cool down. The mass of the white crystals, when cooled was recorded. The experiment compound was safely removed and the materials used were properly stowed away. Table 1: MassesMass of Beaker and Stiring Rod 77.774gMass of Beaker + Hydrated Copper(II) Sulfate 81.531gMass of White Crystals 80.348gMass of CuSO4 x H2O Mass or H2O= 18.0152g Mass of CuSO4= 159.6068g2(18.0152g Mass of CuSO4= 159.6068g2(18,015 1.00794)1(1). 15.9994) Mass of hydrated CuSO4= 3,757g Mass of water-free= 2,574g1(6 3,546) 1(32,065) 4(15.9994)1 177.62g Analysis questions page 227 in Text1a.) 3,757 Difference: 1,183 / 3,757 x 100% = 31.49%-2,574 1,183b.) It is expected that the mass percentage of water should be comparable in all groups, with respect to the Hydrated CuSO4 x H2O.2a.) Table 2: Class Group ResultsGroup Mass(CuSO4 x H2O) Mass(CuSO4) % H2O1 3,757g 2,574g 31.49%2 3,587g 2,277g 36.52%3 3,871g 2,646g 31.65%4 3,188g 2,075g 34.91%5 3,465g 2,168g 37.43%b.) 3.) CuSO4= 100 - 31.49 H2O Class Average= 34.4%= 65.6g34.4g (1 mol/18.01528)=1,909mols65.4g 6g (1 mol/ 159.6086) Therefore the Molecular Formula Is CuSO4 x 5H2O.=0.41mols1.909/ 0.41= 4.66= 5 H2O molecules4.) If you ionic compound in a test tube would heat you would expect steam to rise, which means that the water in the compound is evaporated in the air. In our experiment we saw this happen, but it would be more obvious in a test tube, because it is more limited than the ARE THE QUESTIONS I AM VERY CONFUSSED WITH, YOUR help is much appericiated:1. You obtained the mass percentage of water in the copper sulphate hydrate.a.) Use your obervations to calculate the percentage composition of the copper sulfaathydrate.b.) In the case of a hydrate, and assuming that you know the formula of the corresponding water-free ionic compound, do you think it is more useful to have the mass perventage composition? Explain your answer.2. Let's say you don't compelely convert the hydrate into the water-free compound. Explain how this would affect the calculated percent by mass of water in the compound. and the molecular formula you determined.3. Suppose the hydrate heated too quickly and part of it spewed out of the container and was lost. Explain how this would affect the calculated percentage of the mass of water in the compound and the molecular formula you determined. (I understand it would change, but exactuly how?) 4. Suggest a source of errors (not yet mentioned) that would result in a value of x that isa.) higher than the actual valueb.) lower than the actual value +100Join Yahoo Answers and get 100 points today. Terms • Privacy • AdChoices • RSS • Answers • Community Guidelines • Leaderboard • Knowledge Partners • Points & LevelsSend Feedback> +100Join Yahoo Answers and receive 100 points today. Terms • Privacy • AdChoices • RSS • Answers • Community Guidelines • Leaderboard • Knowledge Partners • Points & LevelsSend Feedback> Feedback> Feedback>

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