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## Stoichiometry of a precipitation reaction lab answers

Copyright 2019 - Hands-on Labs | 1/3 student name Mohammed Alnaimi Student ID 48408 Precipitation Response Agency Ocean County University Session 2019L2 CHEM 181 DL1 Course CHEM 181 DL1 Instructor Leah Stage Final Report Practice 1 CHEM 1CHEM 181 DL1 Stoyciometry's Precipitation ion Reaction 86% It was not optimal 100% were in a higher range. If it hadn't gone to number 10, the yield could have been higher. My theoretical yield was 0.68g. However, since I went to scale only the 10th place, I can't say with certainty that my product weighed only 0.6 g. It may have been 0.68 g but I don't know it. There 1. Given that the theoretical yield for isolating calcium carbonate in this experiment will be 100%, based on the information and the results obtained from this experiment, we explain the success stories in the recovery of calcium carbonate and suggest two sources of error that prevented us from obtaining 100% yield. Excess products of CaCO<sub>3</sub> will be produced because there is not enough CaCl<sub>2</sub> to respond to Na<sub>2</sub>CO<sub>3</sub>. The purpose of stoichiometric quantities is to know the exact amount of each reaction or to be close to the exact amount in order to produce a product that is close to the theoretical amount and therefore has a higher percentage yield. 2. How does it affect the stoichiometric quantity of the quantity produced to add twice as much CO as is required? 2 3 Copyright 2019 - Hands-on Labs | 2 / 3 Data Table 1: Stoichiometry Value 7.5gCaCl<sub>2</sub>·9H<sub>2</sub>O x (1molCaCl<sub>2</sub>·9H<sub>2</sub>O)/273.16gCaCl<sub>2</sub>·9H<sub>2</sub>O)x (1molCaCl<sub>2</sub>/1molCaCl<sub>2</sub>·9H<sub>2</sub>O). At CaCl of 7.5 g, the quantity (g) of pure CaCl is to be 0.5 g .9H O. Shows your work.2 2 4g MgSO<sub>4</sub> ·7H<sub>2</sub>O x (1mol MgSO<sub>4</sub> ·7H<sub>2</sub>O / 246.51g MgSO<sub>4</sub> ·7H<sub>2</sub>O)x (1mol MgSO<sub>4</sub>/1mol MgSO<sub>4</sub>·7H<sub>2</sub>O) At 2.4 g of MgSO<sub>4</sub>, it is possible to decide the quantity (g) of pure MgSO<sub>4</sub> ·7H<sub>2</sub>O. Shows your work.4 4 2 According to the background, the reaction agent of our established total mass should be the same as the total mass of the product in hand. Therefore, we must expect a slightly smaller amount of products, since in real experiments the system is almost completely closed. The preservation of mass can be confirmed in experiments by measuring the weight of the reactant and weighing the product. The total mass of the reactant must be slightly larger than the weight of the product. 5. Mass preservation was discussed in the background. Explains how mass retention (non-theoretical real world) can be examined in experiments. Copyright 2019 - Hands-on Labs | 3/3 Early: CaCl ·2H O (g) Early: CaCl · 2H O (mol) Early: CaCl (mol) Early: Na CO (mol) Early: Na CO (g) Theory: CaCO (g) Filter Paper Mass (g) Filter paper + CaCO (g) Actual: CaCO (g) % Yield: 2 2 2 1.0 g 2 0.0068 mol 2 0.0068 mol 2 3 0.0 068 mol 2 3 0.8g 3 0.68g 0.9g 3 1.5g 3 0.6g 86% © 2020 Sob | Solutions for professional online tutoring help

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