100 greatest discoveries chemistry worksheet

What are the most important scientific discoveries of all time? The answer may include Copernicus's idea that the sun is at the heart of the solar system, Newton's laws of motion, the Mendele's periodic table of elements, Einstein's theory of relativity and Darwin's theory of evolution. There are a number of great scientific discoveries that have had a significant impact on the way we think and live through history. The 100 Greatest Discoveries recognises the 100 most important scientific discoveries of all time and explains them with historical re-creations, archival imagery and interviews with scientists. 100 Greatest Discoveries - Chemistry This episode recognises thirteen important discoveries in chemistry, including the discovery of oxygen by Joseph Priestley, the atomic theory of John Dalton and the discovery of the periodic table of Dmitry Mendeleev's elements. Here are thirteen important discoveries: 1. Oxygen (1770)2. Atomic Theory (1808) 3. Atoms merge into molecules (1811 onwards) 4. Urea Synthesis (1828) 5. Chemical structure (1850)6. Periodic table of elements (1860-1870) 7. Electricity Transforms Chemicals (1807 - 1810) 8. Electron (1897) 9. Electrons for Chemical Bonds (1913 onwards) 10. Atoms have signatures of light (1850)11. Radioactivity (1890-1900) 12. Plastics (1869 and 1900) 13. Fullerenes (1985) Scientists have changed the way we think and live through the centuries. What are the most important scientific discoveries of all time? In no particular order we present the top 100 in eight different categories. 100 Greatest Discoveries - CHEMISTRY 1. Oxygen (1770)Joseph Priestley discovers oxygen; later Antoine Lavoisier explains the nature of the elements. Priestley produces oxygen in experiments and describes his role in burning and breathing. Then, by dissolving the fixed air in the water, it invents the water. Priestley, regardless of the importance of his discovery, calls the new gas dephlogized air. Lavoisier gives oxygen his name and correctly describes his role in combustion. Lavoisier then collaborates with others in the design of the chemical nomenclature, which serves as the basis of the modern system. 2. Atomic Theory (1808) John Dalton provides a way of connecting invisible atoms with reconnaissance quantities, such as gas volume or mineral mass. Its atomic theory indicates that the elements are composed of tiny particles called atoms. Thus, the pure element consists of identical atoms, all of them of the same mass, and the compounds consist of atoms of different elements grouped together. 3. Atoms merge into molecules (1811 onwards) Italian chemist Amedeo Avogadro finds that atoms in the elements combine and form molecules. Avogadro suggests that the same amounts of gases under the same temperature and pressure conditions contain the same number of molecules. 4. Urea synthesis (1828) Friedrich Woehler accidentally synthesizes urea from substances proving that substances produced by living beings may be repless by non-living substances. Until 1828, it was believed that organic matter could only be formed by the vital force present in animals and plants. 5. Chemical structure (1850) Friedrich Kekule finds the chemical structure of benzene, which brings a study of molecular structure to the forehead chemistry. It says that after years of studying the nature of carbon-carbon bonds, he came to the ring-shaped molecule benzene after dreaming that the snake would seize its tail. The unusual structure solves the problem of how carbon atoms can connect to up to four other atoms at once. 6. Periodic table of elements (1860-1870) Dmitry Mendeleyev is aware that if all of the 63 known elements are arranged in order of increase in atomic weight, their properties are repeated according to certain periodic cycles. It designs a periodic table of elements and predicts the existence of elements that have not yet been detected. Three of these elements are found during his life: gallium, candied and Germanium. 7. Electricity Transforms Chemicals (1807 – 1810) Humphry Davy finds that electricity transformis chemicals. It uses an electric heap (an early battery) to separate salts by a process now known as electrolysis. With numerous batteries it can distinguish elemental potassium and sodium in calcium, strontium, barium and magnesium. 8. Electron (1897) J.J. Thomson discovers that the negatively charged particles emitted by cathode ray tubes are smaller than atoms and part of all atoms. These particles, now known as electrons, are called corps. 9. Electrons for Chemical Bonds (1913 onwards) Niels Bohr publishes its model of the atomic structure in which electrons travel in specific orbits around the ailing surface, and the chemical properties of the element are largely determined by the number of electrons in its output orbits. This connects the way to understanding how electrons are involved in chemical binder. 10. Atoms have signatures of light (1850)Gustav Kirchhoff and Robert Bunsen find that each element absorbs or emits light at certain wavelengths and produces specific spectra. 11. Radioactivity (1890-1900)Marie and Pierre Curie detect and isolate radioactive materials. After chemical extraction of uranium from uranium, Marie notes that the remaining material is more active than pure uranium. He notes that there are new elements in the traces, in addition to uranium, that are radioactive. This leads to the discovery of elements of polonium and radium. 12. Plastics (1869 and 1900)John Wesley Hyatt formulates celluloid plastics for use as a substitute for ivory in the production of billiard balls. Celluloid is the first significant synthetic plastic and is used as a substitute for expensive substances such as ivory, amber, horn and turtle. Later Leo Baekeland invents tempered plastic, specifically Bakelite, a synthetic substitute for shellac used in electronic insulation. 13. Fullerenes (1985)Robert Curl, Harold Kroto and Rick Smalley are discovering a whole new class of carbon compounds with a cage-like structure. This leads to the discovery of similar tube-like carbon structures. Together, the compounds are called buckminsterfullerenes, or fullerenes. The molecules are made up entirely of carbon and are in the form of a hollow sphere, ellipsoid, tube or ring. Named after Richard Buckminster Fuller, the architect who created the geodetic dome, they are sometimes called buckyballs or buckytubes. Thanks for the cooperation! A student worksheet that needs to be completed while watching a video. The video is available on YouTube. Worksheet includes:One page with 15 questions enteredReans the Go TPT credit key for future purchases! Go to the My Purchases page. Next to each purchase, click the Provide feedback button. Give a review and leave a comment on the product. When you give feedback, TPT gives you credits that can reduce the cost of future purchases. Rate this product and earn TPT CREDIT! Click the green star under the name of this store and Follow GertrudeKatzChronicleTo you will be the first to know about sales, freebies and product launches! StandardsAnual meaning of symbols, keywords, and other domain words and phrases, as they are used in a specific scientific or technical context relevant to 9-10 rating texts and topics. Themes.

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