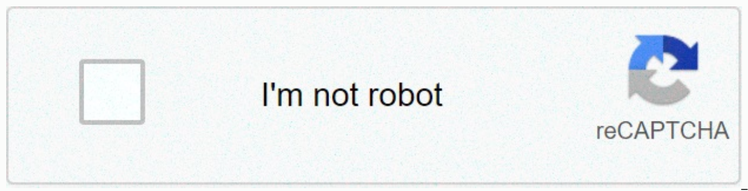


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# Chemistry formula sheet 2019

Sorry, you've disabled Javascript! To see this page as it is meant to appear, turn on your Javascript! See instructions here Sorry, you've disabled Javascript! To see this page as it is meant to appear, turn on your Javascript! See instructions here by John T. Moore During the first year of high school chemistry or the first semester of college chemistry, all terms, units of measurement, and atoms, molecules, elements and compounds may seem a little overwhelming. The good news: Your Chem I class doesn't have to be torture. This handy Cheat Sheet provides some basic information that you refer to regularly to make your chemistry class just a little bit easier. When attached, atoms lose, win or divide to have the same number of electrons as the precious gas closest to the periodic table. Ionic, covalent and metal bonds are formed by combinations of metals and nonmetals. Metal + nonmetal = ionic binding Nonmetal + nonmetal = co-deposits binding Metal + metal = metal binding When two elements engage in ionic binding, one or more electrons are transferred from the metal to the non-metal, forming ions (charged atoms). The metal, which has lost one or more electrons, forms a cation, an ion with a positive charge; the non-metal, which has extracted one or more electrons, becomes an anion, an ion with a negative charge. When two elements form a covalent bond, one or more electron pairs are shared between these two elements. In metal bonding, which occurs in metals (either a pure metal or an alloy of two or more metals), the valence (outer shell) electrons are donated to a sea of electrons. Much chemistry is explained by the sharing and trading of electrons between atoms. Understanding how electrons are arranged in an atom is a building block of Chem I. Electrons in an atom are incorporated into specific energy levels (1, 2, 3, and so on) representing different distances from the nucleus. The greater the energy level, the further it is from the core. Electrons that are at the highest energy level are called valence electrons. Within each energy level there is a space volume where specific electrons are likely to be located. These spaces, called orbitals, are of various forms, indicated by a letter (s, p, d, f, g). (In most cases, only the electrons in the s- and p-orbitals are considered valence electrons.) Electrons are looking for the lowest possible energy level. The following electron-filling pattern indicates how the electrons fill up in the energy levels. Knowing this pattern is useful in many aspects of chemistry, including predict the binding situation of a particular atom and in the prediction of the geometry of a covalent compound. Electron filler pattern: 1s, 2s, 2p, 3p, 3p, 4s, 3d, 4p, 5s, 4d, 5p, 6p, 5p, 6p, 7s, 5f The mole (abbreviated mole and also called Avogadro's number) is a conversion conversion that enables a chemist or chemistry student to move from the microscopic world of atoms, ions and molecules to the macroscopic world of grams, kilograms and tons. The mole conversion is used in reaction stoichiometry to predict how much product can be made from a certain amount of reactant or how much reactant is needed to produce a certain amount of product. If you know the particles, moles or grams of a substance, you calculate the other two measurements using the following equation: 1 mole = 6,022 × 1023 particles/mole = formula weight expressed in grams in Chem I, you may need to identify isotopes, which are atoms of the same element that have different numbers of neutrons. The following view identifies a specific isotope of an element. It is widely used in balancing nuclear reactions. X = element symbol Z = atomic number (number of protons) A = mass number (number of protons + number of neutrons) To succeed in your Chem I class, you need to have a solid understanding of basic chemistry measurements and how to convert them from one measurement to another. Below are some important conversions of temperature, size, and pressure as well as metric prefixes to remember for your chemistry class: Temperature conversions: °F = 9/5(°C) + 32 °C = 5/9(°F – 32) K = °C + 273 English/metric conversions: 1 in = 2.54 cm 1 lb = 454 g 1 qt = 0,946 L Pressure conversion: 1atm = 760 mmHg = 760 torr Ordinary metric prefixes: milli- = 0,001 centi- = 0.01 kilo- = 1,000 Much chemistry requires you to understand the difference between acids and bases. An acid is a substance that donates a H+ ion to another chemical called a base. A base is a substance that accepts (combines with) an H+ ion. If you want to know the concentration of the H+ ion in solution, you do so by representing the H+ molarity [H+]. Another way to display the H+ concentration is the pH, which is the negative logarithm of the H+ molarity. The following equation shows this mathematical relationship and the way to calculate H+ molarity, given the pH: pH = – log[H+]; [H+] = 10-pH pH = 7 is neutral. pH less than 7 is acidic. pH larger than 7 is basic. When studying the properties of gases, you need to know the relationships between the variables of volume (V), pressure (P), Kelvin temperature (T) and the amount in moles (n), so that you can calculate missing information (P, V, T, or n) and solve reaction problems. Although the pairs of variables have individual relationships, the two most important and useful gas laws are the combined gas law and the ideal gas law: Combined gas law (P1V1)/T1 = (P2V2)/T2 (T must in Kelvin) Ideal gas law PV = nRT (R = 0.0821 L atm/K.mol) Show: All subjects See also Pūtaiao Curriculum & Standards documents Standards 1 standards Level 2 standards Level 3 Standards Level 3 3 Chemistry Matrices NCEA at TKI Chemistry Chemistry Remote Assessment Matrix and Guidance (PDF, 150KB) Education and Learning Guide TKI All levels Resources for internally assessed standards Clarifications All levels Examples of student work All levels Assessment Evidence Gathering Templates All levels National Moderator's Report All levels TKI Resources & Conditions of Assessment Level 1 Level 2 Level 3 Sign in to your Facebook account to view the NZQA Chemistry Facebook page Click here to browse spotlights resources for externally rated standards Level 1 Level 2 Level 3 Assessment Assessment Specifications Evidence Gathering Templates All levels Exam papers and copies Level 1 Level 2 Level 3 Assessment Schedules Level 1 Level 2 Level 3 Cut Scores Level 1 Level 2 Level 3 Level 3 Assessment Level 1 Level 2 Level 3 Exam Papers for Expired Standards These sources were removed , because the standards have expired and have been replaced by new standards and resources. 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