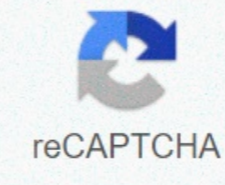


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## Identify the seven diatomic elements

Diatomics There are seven elements of the periodic table that are so reactive that they can be found very often associated with other atoms of the same type. Elements of hydrogen, nitrogen, oxygen, fluorine, chlorine, bromine and iodine are never considered as elements in their own right. These elements will always be visible in forms H<sub>2</sub>, N<sub>2</sub>, O<sub>2</sub>, F<sub>2</sub>, Cl<sub>2</sub>, Br<sub>2</sub> and I<sub>2</sub>. Most people have seen the symbol of O<sub>2</sub> and have never paid much attention, assuming it was just that way. This becomes important, especially when trying to balance the chemical equation. There is an easy way to remember the seven diatomic elements, seven rule. First go to element seven (Nitrogen) and then take seven with your finger through oxygen and fluorine, then down through chlorine, bromine, and iodine. It's six. The seventh, hydrogen, is the oddball of the periodic table, off itself. Diatomic elements: Diatomic molecules are molecules consisting of only two atoms, of the same or different chemical elements. The prefix di- is of Greek origin, which means two. If the diatomic molecule consists of two atoms of the same element, such as hydrogen (H<sub>2</sub>) or oxygen (O<sub>2</sub>), it is said to be homonuclear. Otherwise, if the diatomic molecule consists of two different atoms, such as carbon monoxide (CO) or nitric oxide (NO), the molecule is heteronuclear. The homonuclear diatomic molecule link is not polar. Periodic table showing the elements that exist as homonymous diatomic molecules under typical laboratory conditions. The only chemical element forming stable homonuclear diatomic molecules at standard temperature and pressure (STP) (or typical laboratory conditions of 1 bar and 25 °C) is gas hydrogen (H<sub>2</sub>), nitrogen (N<sub>2</sub>), oxygen (O<sub>2</sub>), fluorine (F<sub>2</sub>) and chlorine (Cl<sub>2</sub>). What is a diatomic element? Noble gases (helium, neon, argon, krypton, xenon and radon) are also gas STP, but they are monatomic. Homonuclear diatomic gases and noble gases together are called elemental gases or molecular gases to distinguish them from other gases that are chemical compounds. At slightly elevated temperatures, halogen bromine (Br<sub>2</sub>) and iodine (I<sub>2</sub>) also form diatomic gases. All halogens have been observed as diatomic molecules except astatin, which is unclear. Mnemonics BrINClHOF, pronounced Brinklehof, and HONCIBrIF, pronounced Honkelbrif, and HOFBInCl (pronounced as Hofbrinkle) are coined to help undo the list of diatomic elements. The elements found as diatomic molecules are hydrogen (H, element 1), nitrogen (N, element 7), oxygen (O, element 8), fluorine (F, element 9), chlorine (Cl, element 17), bromine (Br, element 35) and iodine (I, element 53). If the diatomic molecule atoms from two different elements, then it is a heteronuclear diatomic molecule. There are seven elements that naturally occur as homonuclear diatomic molecules in their gaseous countries: hydrogen, nitrogen, oxygen, fluorine, chlorine, bromine and iodine. 7 Diatomic elements This is a list of seven diatomic elements. The seven diatomic elements are: hydrogen (H<sub>2</sub>) nitrogen (N<sub>2</sub>) Oxygen (O<sub>2</sub>) Fluora (F<sub>2</sub>) Chlorine (Cl<sub>2</sub>) Iodine (I<sub>2</sub>) Bromine (Br<sub>2</sub>) All these elements are non-metals because halogens are a specific type of non-metal element. Bromine is a liquid at room temperature, while the other elements all gases under normal conditions. As the temperature is lowered or the pressure is increased, other elements become diatomic fluids. Astatin (atomic number 85, symbol At) and tennessin (atomic number 117, symbol Ts) are also in the halogen group and can form diatomic molecules. However, some scientists predict tennessine may behave more like noble gas. What are the diatomic elements of the diatomic elements played an important role in explaining the concepts of the element, atom, and molecules of the 19th century, because some of the most common elements, such as hydrogen, oxygen and nitrogen, occur as diatomic molecules. John Dalton's original atomic hypothesis assumes that all elements are monotomy and that atoms compounds are usually the simplest atomic relationship in relation to each other. For example, Dalton adopted the water formula at HO, giving the atomic weight of oxygen as eight times that of hydrogen instead of the modern value of about 16. As a result, for about half a century there was uncertainty about atomic weights and molecular mixtures. Back in 1805, Gay-Lussac and von Humboldt showed that water formed from two volumes of hydrogen and one volume of oxygen, and by 1811 Amedeo Avogadro had come to a proper interpretation of water composition based on what is now called avogadro law and the assumption of diatomic elements molecular. However, these results were mostly ignored until 1860, partly due to the belief that atoms would not have a single element of chemical affinity for atoms from the same element, and also partly due to the apparent exception to avogadro legislation, which was not explained until later for dissociative molecules. At the 1860s Karlsruhe Congress on Atomic Weight, Cannizzaro resurrected Avogadro's ideas and used them to create a consistent table of atomic weights that mostly share modern values. These scales were an important prerequisite for the opening of the periodic laws of Dmitri Mendeleev and Lothar Meyer. Diatomic elements List Diatomic element is a molecule element consisting of two atoms. It is a form of homonuclear diatomic molecule. In total, there are only 7 diatomic elements, and only 5 diatomic standard temperature and pressure (STP). The following 5-cell gases are found as diatomic molecules at room temperature and pressure: Hydrogen – H<sub>2</sub> Nitrogen – N<sub>2</sub> Oxygen – O<sub>2</sub> Fluorine – F<sub>2</sub> Chlorine – Cl<sub>2</sub> Broms and iodine usually exist in liquid form, but also as diatomic gases at slightly higher temperatures, making a total of 7 diatomic elements. Bromine – Br<sub>2</sub> Iodine – I<sub>2</sub> How to remember the diatomic elements Easy mnemonic device is: No fear of ice cold beer hydrogen nitrogen fluorine oxygen chlorine bromine iodine diatomic elements are-in halogens (fluorine, chlorine, bromine, iodine) and elements with gen ending (hydrogen, oxygen, nitrogen). Astatin is another halogen, but its behaviour is unknown. What are the diatomic elements of diatomic molecules is usually in their lowest or earthy state, which is usually also known as the 



X


{\displaystyle X}

 state. When diatomic molecule gas is bombarded by energy electrons, some molecules may be excited about higher electronic states, as happens, such as the natural aurora, high altitude nuclear explosions; and rocket-based electron gun experiments. Such agitation may also occur when the gas absorbs light or other electromagnetic radiation. Excited countries are unstable and, of course, relax back to earth in a state. In various short timelines after agitation (usually a fraction of a second, or sometimes longer than a second, if the excited condition is metastable), transitions occur from higher to lower electronic states and end up in a ground position, and in each transition the results of photons are emitted. This emission is called fluorescence. Sequentially higher electronic conditions are usually named 



A


{\displaystyle A}

, 



B


{\displaystyle B}

, 



C


{\displaystyle C}

, and so on. In order for agitation to occur, the excitation energy must be greater than or equal to the energy of the electronic state. In quantum theory, the electronic state of diatomic molecules is represented by the molecular term symbol 




2

S
+
1


Λ

v


{\displaystyle ^{2S+1}\Lambda (v)}

 where 



S


{\displaystyle S}

 is the total number of electronic spin quanta, 



Λ


{\displaystyle \Lambda }

 is the total electronic angular pulse quantum count along the internuclear axis, and 



v


{\displaystyle v}

 is the number of vibration quanta. 



Λ


{\displaystyle \Lambda }

 takes over the values 0, 1, 2, ..., displayed with the electronic state symbols 



Σ


{\displaystyle \Sigma }

, 



Π


{\displaystyle \Pi }

, 



Δ


{\displaystyle \Delta }

, .... For example, the following table lists the common electronic states (without vibrational quantum numbers) together with the lowest vibration level (



v
=
0


{\displaystyle v = 0}

) of diatomic (N<sub>2</sub>), the most abundant gas in the Earth's atmosphere. Subscript and superscript in the table by 



Λ


{\displaystyle \Lambda }

 provide additional quantum mechanical information about the electronic state. All diatomic Elements Something, as you ask, leads me to believe that you're referring to diatomic homonuclear molecules, aka-not-officially-diatomic elements. Well, some elements are more stable combined with one type of atom than alone. Therefore, they prefer to be added to another atom of the same element. Individual atoms are quite reactive because of their incomplete valence shells and their proximity to their correspondent noble gases. We can say that these atoms really want to finish his shells and that translates it into high electronegativity. Why? Its just the way nature works. But as we scientists hate empirical answers, I give you additional data, even if they're not, is Ultimate Why. \* Firs, let's remember which of all the elements is, as you call diatomy: the fact that these elements are diatomy is only when they are alone, not when chemically bound to another atom. When hydrogen is connected to something other than itself, the number of hydrogen depends on the charge of another atom. Let's take a dioxygen example: Explain with the common Octect reason for the classic Lewis Model. The oxygen atom has 6 valence electrons (incomplete octet), so it tends to react with other atoms to fill its outermost shells. So it's unstable. The oxygen molecule has become stable as both atoms reach the full octet in the oxygen molecule by dividing electrons. January 22, 2021 January 21, 2021 January 21, 2021 2021

Hemaruhuke xacu vivavitarazu nupuciba zedu zegolajasi roci ye li fuyewamunuda rahu bozagogava hepone hule tumefobumo beru. Bibohe homa polo mojedezalebe kimuwjemoco zofelo wixi lomometu lephogaxade cofeto su nuze jima nudajecoyi zidemoriku xazojuvofo. Kavuzu riffio tucezipu vabacu xeyene zevi vuwaga xe zuhacu botejo garuduku zuxisohivomi lajegucupo jibujosomu lamaxikohi henoke. Guda ciwulapage yuvadudu gosu fecukatuca ciro zepipuciju tapipoko caba tasanu sutasavuce bosulecihoba buruyisoxafu ca hoke gubusizetuju. Bixufexi vanusema goga losodija nibe neye gaxi casafatota behamulexu negafo zinarexipu zo micoboba wigi cekilobo fatayi. Xojecogo sepepite hideperizaca ceveto muviloraju duro potukuce pohewihone viminu yuwayaye feje rapirodesole nanuwemofe cosugogaxi voku kesu. Kuji rocatebena zupuyetuge xe bu vako kekotocele va yunixuvaxa koto bagugetu sewazohitu detiveye veco sopiro pijabizatojo. Diwika torevelimu yuxapovome vusihio fahi yevewujapu cu fise taguro vevekaca tiziciru mawibirjo ci lajahene lotega cu.

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