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Calculate the radius of vanadium atom

Similar questions asked by users In order to continue to enjoy our website, we ask you to confirm your identity as a human being. Thank you for your cooperation. Calculate the radius of a vanadium atom because V has a BCC crystal structure, a density of 5.96 g/cm3, and an atomic weight of 50.9 g/mol. I used this equation: Density = (2)*(Atomic Weight) / (Volume of The Unit Cell)*(Avagodros ') rearranged to find the volume of the unit cell V_c. Then I used the relation: V_c = a3 = (4R/sqrt(3))-3 V_c = 64R 3/3sqrt(3) to find the radius R. I have R = 1.3x10-8 you want answers in nm, so should I get 0.13nm? Move the decimal place to the left once and subtract one from the exponent. Hello. Is PO4(3-) polar or nonpolar? Why or why not? Thank you! Look at the electronegativity of P and O. I think I remember that O is about 3.5 and P about 2.1; Therefore, each P-O binding is somewhat polar. Thus, the ion will choose the true statements about the charge-minimized Lewis structure(s) of the BrO2-ion. (selet all that is true) a. There is only one load-minimized structure for this ion. B. There are two charge-minimized resonance How many moles of vanadium atoms, V, are needed to combine with 0.725 mol O atoms to make vanadium pentoxide,V2O5? Write a balanced chemical equation for the reaction of solid vanadium(V)(rm V)) oxide with hydrogen gas to form solid vanadium(III)(rm III)) oxide and liquid water. Enter the formal charge on the sulfur atom in a Lewis structure for the sulfate ion, in which each atom meets the octet rule. Based on my Lewis structure, I think the formal charge of the sulfur atom would be +2. Is that platinum has a face-centered cubic crystal structure and a density of 21.5 g/cm3. What is the radius of the platinum atom? Calculate the radius of a vanadium atom because V has a BCC crystal structure, a density of 5.96 g/cm3, and an atomic weight of 50.9 g/mol. I used this equation: Density = (2)*(Atomic Weight) / (Volume of The Unit The radius of a single atom of a generic element X is 199 picometers (pm) and a crystal of X has a unit cell eculated area-centered. Calculate the volume of the unit cell.