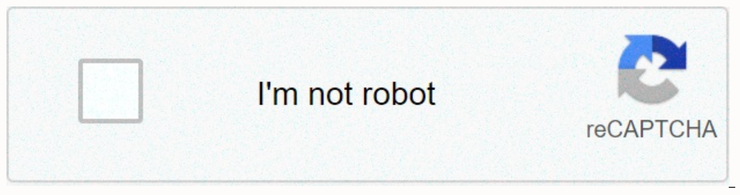

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Endothermic reaction has positive delta h

+100Join Yahoo Answers and get 100 points today. Terms • Privacy • AdChoices • RSS • Reply Help • Community Guidelines • Leaderboard • Knowledge Partners • Body & LevelsSend Feedback • Moderators: Chem_Mod, Chem_Admin rachana_korapati3C Posts: 39 Joined: Fri Sep 28, 2018 12:25 Pm Postby rachana_korapati3C » Tue Jan 15, 2019 16:36 I don't think you can tell if the reaction is exothermic or endothermic by the equation itself. You either need to measure heating for cooling, or know enough about similar types of reactions. Nathan Tran 4K Posts: 92 Joined: Fri Sep 28, 2018 12:16 Pm Postby Nathan Tran 4K » Tue Jan 15, 2019 17:34 I know that the triangle at the top of the arrow indicates that heat is involved, so you know that the equation can be endothermic or exothermic. However, to determine whether the reaction is exothermic or endothermic, I assume we'd get delta H or some variant of it. juliasloan_4g Posts: 67 Joined: Fri Sep 28, 2018 12:28 Pm Postby juliasloan_4g » Tue Jan 15, 2019 18:19 Yes, you should get delta h to see if the action is endo(at heat) or exo (giving away heat) thermal jonathanjchang2E Contributions: 61 Linked: Fri 28 September, 2018 12:26 Pm Postby jonathanjchang2E » Tue Jan 15, 2019 18:31 You can predict whether the reaction is endothermic or exothermic when looking at the energy of reactants and products (but this way it's really hard with more complex reactions). For example, if you get H2 = 2H, it would probably be endothermic because H2 is less energy than 2 H. shaunajava2e Contributions: 66 Joined: Fri Sep 28, 2018 12:26 Pm Postby shaunajava2e » Wed Jan 16, 2019 10:22 am Rapid difference between endothermic and exothermic includes reactions in the environment. Endothermic reaction occurs when energy is

absorbed from the surrounding area in the form of heat, and exothermic is when energy is released from the system into the surrounding area. Samantha Chang 2K Posts: 69 Joined: Fri Sep 28, 2018 12:17 Am Postby Samantha Chang 2K » Wed Jan 16, 2019 12:43 pm We can know if the reaction is endothermic when heat absorbs and vice versa exothermic reaction is when heat is released. This is how one can determine how it is possible to determine whether the reaction is exothermic and endothermic. Besides, I think you need to know if Delta H is negative or positive. Melissa Bu 1B Posts: 36 Joined: Fri September 28, 2018 12:19 Postby Melissa Bu 1B » Wed Jan 16, 2019 14:45 Helps to see the energy graph to think about why the exothermic has a negative change in H and endothermic has a positive change in H. For exothermic reactions, the system leaves heat, so the graph ends up in products at a lower delta H value compared to reactants. Therefore, the change in V H will be negative (smaller H - greater H < 0). In endothermic reactions, heat is inserted into the system/the system is absorbed to hear, so that the higher delta H compared to reactants. Therefore, the change in V H will be positive (greater H - smaller H > 0). Cynthia Aragon 1B Posts: 47 Connected: After April 09, 2018 1:38 pm Postby Cynthia Aragon 1B » Wed Jan 16, 2019 14:52 Endothermic reaction absorbs energy from ambient and exothermic reactions releases/produces heat. Endotermic: delta H >0Exothermic: delta H < 0 Nicolle Fernandez 1E Posts: 56 United: Fri Sep 28, 2018 12:25 Elected: 1 time Postby Nicolle Fernandez 1E » Wed Jan 16, 2019 15:58 When the reaction emits heat, it is exothermic, but when the reaction requires heat, it is endothermic. THE REACTION YOU MUST BE EXOTHERMIC, change H must be less than 0. In order for the reaction to be endothermic, the H change is greater than 0. Go back to Using Le Chatelier principle changes in chemical & physical conditions Go to the user browsing this forum: No registered users and 0 guests your original equation is incorrect. For a process that occurs in a closed system, the equation should read $$\Delta U=Q+W$$where U is the internal energy of the system, Q is the heat added to the system, and W is the work done by the environment in the system, $W=-\int{PdV}$. If the process is running at constant pressure, $W=-P\Delta V$, and the change in internal energy becomes:$$\Delta U=Q-P\Delta V$$. But by definition enthalpy, we have $\Delta H=\Delta U+\Delta (PV)$. So, finally, $$\Delta H=Q$$ So for a process performed at constant pressure, if the heat added to the system is positive (endothermic), ΔH is positive, and if the heat added to the system is negative (exothermic, heat removed from the system), ΔH is negative. Copyright © 2020 Multiply Media, LLC. All rights reserved. The material on this site may not be reproduced, distributed, transmitted, cached or otherwise used, except with multiply's prior written consent. Ohio State University University

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