Chemistry chapter 18 assessment answers

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1. The al-Black metals have a single S-electric in their outer shell. In contrast, alkaline earth metals have a complete s subshell in their outer shell. Generally, alkaline metals react faster and are more reacting to the related alkaline earth metals in the same period. 3. Na + I2 ⟶ 2NaI2Na + Se ⟶ Na2Se2Na + O2 ⟶ Na2O2Na + I2 ⟶ 2NaI2Na + Se ⟶ Na2Se2Na + O2 ⟶ Na2O2 Sr + I2 ⟶ SrI2Sr + Se ⟶ SrSe2Sr + O2 ⟶ 2SrOSr + I2 ⟶ SrI2Sr + Se ⟶ SrSe2Sr + O2 ⟶ 2S⟶ 2Al + 3I2 2Srosr + 3Se ⟶ Al2Se34Al + 3I2 ⟶ 2Al4o32al + 3I2 ⟶ 2Srosr + 3Se ⟶ Al2Se34Al + 3I2 ⟶ 2Al2O35. Possible methods of difference between the two include more spectrum than known, a flame test provides yellow color characteristic for sodium (the strontium is a red flame), or compared to their selobalatis in water. At 20 s, NaCl is 35.7 g100 ml compared to 35.7 g100 ml compared to 53.8 ml for G100 with 53.8 ml g100 srCl2 ml. Heat 100 sc provides an easy test, because NaCl's sertra is 39.12 g100 ml, 39.12 g100 ml, but SrCl2 is 100.8 g100 ml. Determining a solid density is sometimes difficult, but there is quite a difference (2.165 g/mL NaCl and 3.052 g/ml SrCl2) that this method will be viable and probably the easiest and least expensive to perform the test. 7. (a) 2Sr (□) + O2 (g) ⟶ 2S2 ;(□) گSr (□) + O2 (g) ⟶ 2S□) گ); (b) Sr (a) sr (a) sr (a) srBr2 (a) sr2 (g); Sr (a) sr (a) sr (a) srBr2 (a) sr2 (g); (g) Sr (a) sr (a) sr (a) srH2 (g) srH2 (a); Sr (a) sr (a) sr h2 (g) srH2 (a); (d) 6Sr (s) + P4 (2 ⟶ (□Sr3P2 (6 ؛ (□Sr (□) + P4 (2 ⟶ (□Sr3P2 (□); (e) Sr (2 + (□H2O (l) ⟶ Sr (□العاقبة) 2 (□او) + H2 (g) Sr (2 + (□H2O (l) ⟶ Sr (□العاقبة) 2 (□او) + H2 (g) 11. Yes, ton of hydrogen responds with hydro-hydroc acid to produce gas. 13. In PbCl2, the relationship is ionic, as indicated by this melting point 501 s c. In PbCl4, the relationship is co-contained, because it is evidenced by the unstable liquid at room temperature. 15.2CsCl (ابل) + Ca (g) → countercurrentfractionatingtower2Cs (g) + CaCl2 (l) 2CsCl (l) + Ca (g) → countercurrentfractionatingtower2Cs (g) + CaCl2 (l) 17. 2 :(كمي) كاكيتهڈو ناKi + + 2Ii − ⟶ 2Li (l) 2 ؛Li + + 2Ii − ⟶ 2Li (l); 2 :(آكسائڈريشن) انوڈICl − ⟶ Cl2 (g) + 2cl −; 2Cl − ⟶ Cl2 (g) + 2cl. 2 :ناعمجموعی ردعمل + + 2Li − ⟶ 2Li (l) + Cl2 (2 ناجلی + + 2Li − ⟶ 2Li (l) + Cl2 (g) 21. Despite its inhibit, magnesium can be used in the construction even when magnesium is going to come into contact with a flame because a protective oxide coating is established, preventing overall oxide. Only if the metal is in a fine technique or is present in a thin sheet will cause its rapid burn. 23. Extract from: Ala (Oh) (a) ((a) : : :