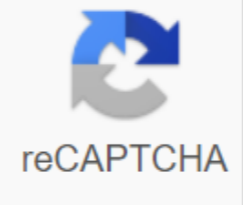




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Alloy type: BinaryContent: Copper and zincDensity: 8.3-8.7 g/cm3Melting Point: 1652-1724 F (900-940 C)Moh's Hardness: 3-4 The exact properties of various brass depend on the composition of brass alloy, in particular, the ratio of copper-zinc. In general, however, all brass are valued for their handling or ease with which metal can be formed into the desired shapes and shapes while maintaining high strength. Although there are differences between brass with high and low zinc content, all brass are considered malleable and ductile (low zinc brass even more so). Because of the low melting point, brass can also be cast relatively easily. However, for casting applications, high zinc content is usually preferable. Brass with lower zinc content can be easily cold, welded and stewed. The high copper content also allows the metal to form a protective layer of oxide (patina) on its surface, which protects against further corrosion, a valuable property in that expose the metal to moisture and weathering. Metal has both good thermal and electrical conductivity (its electrical conductivity can be from 23% to 44%, as is pure copper), and this is this and spark resilience. Like copper, its bacteriostatic properties have led to its use of bathroom fixtures and medical facilities. Brass is considered low friction and non-magnetic alloy, while its acoustic properties have led to its use in many brass orchestra musical instruments. Artists and architects appreciate the aesthetic properties of metal, as it can be produced in a variety of colors, from deep red to golden yellow. The valuable properties of brass and the relative simplicity of production have made it one of the most widely used alloys. Compiling a complete list of all brass applications would be a colossal task, but to get an idea of the industry and the types of products in which brass is located we can categorize and summarize some of the end use based on the class brass used: Free cutting brass (for example, C38500 or 60/40 brass: Nuts, Bolts, threaded partsTerminalsJetsTapsInjectors Copper-zinc alloys were produced as early as the 5th century BC in China and were widely used in Central Asia in the 2nd and 3rd centuries BC These decorative metal pieces, however, are best called natural alloys, as there is no evidence that their producers deliberately alloyed copper and zinc. Instead, it is likely that the alloys were melted from zinc-rich copper ores producing raw brass metals. Greek and Roman documents show that the deliberate production of alloys similar to modern brass, using copper and zinc-rich ore known as calamine, occurred around 1st century BC Kalamina brass was produced using a cement process, resulting in copper being melted in the crucible with the ground smithsonite (or kalamine) ore. At high temperatures, zinc present in such ore turns into steam and permeates copper, thus producing relatively pure brass with a content of 17-30% zinc. This method of brass production was used for almost 2,000 years until the early 19th century. Shortly after the Romans discovered how to produce brass, the alloy was used to mint coins in areas of modern Turkey. It soon spread throughout the Roman Empire. Brass is a generic term that refers to a wide range of copper-zinc alloys. In fact, there are over 60 different types of brass specified in (European norm) standards. These alloys can have a wide range of different compositions depending on the properties required for a particular application. Brass is most often made from copper scrap and zinc bars. Scrap copper is selected based on its impurities, as some additional elements are desirable in order to produce the exact variety of brass required. As zinc begins to boil and evaporate at 1665 degrees Fahrenheit (907 degrees Celsius), below the melting point of copper 1981 degrees Fahrenheit (1083 degrees Celsius), copper must be melted. Once melted, zinc is added to the ratio corresponding to the class of brass produced. Although some manuals are still made for the loss of zinc for evaporation. At the same time, any other additional metals, such as lead, aluminum, silicon or arsenic, are added to the mixture to create the desired alloy. Once the molten alloy is ready, it is poured into molds where it hardens into large slabs or blanks. The blanks - most often alpha beta brass - can be directly processed into wires, pipes and tubes using a hot extrusion that involves pressing the heated metal through a melting pot, or a hot forging. If not extruded or forged, the blanks are then heated and fed through steel rollers (a process known as hot rolling). The result is slabs less than half an inch thick (13 mm). After cooling the brass is fed through the milling machine, or scalper, which cuts out a thin layer of metal to remove defects in the casting surface and oxide. Under the gas atmosphere to prevent oxidation, the alloy is heated and rolled again, a process known as annealing, before it is rolled again at lower temperatures (cold rolling) on sheets about 0.1 (2.5 mm) thick. The cold rolling process deforms the inner structure of the brass grain, resulting in a much stronger and harder metal. This step can be repeated until the desired thickness or hardness is achieved. Finally, the sheets are sawing and haircuts to get the width and length required. All sheets, cast, forged and extruded brass materials are given to a chemical bath, usually one of salt and sulphuric acid, to remove black copper oxide scale and tarnish. Tarnish, metals and properties pdf, properties of metals and properties of nonmetals, transition metals and properties, uses of metals and properties, list of metals and properties, non metals and properties, alkali metals and properties, ferrous metals and properties

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