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This is a measure of how much things change over a period of time. How can you express the speed of the chemical reaction? Changing the number of product reactionians per unit of time. Identify the activation energy. The minimum energy that encounters particles must have in order to react. List four factors that can affect the rate of chemical reaction. - Temperature-Concentration- Particle Size/Excitement- Catalysts What is Chemical Balance? This is when the speed forward reaction and the feedback is equal. Balance review pg. 610 What happens to the amount of reactionary means and products after the reaction has reached the chemical equilibrium. Since chemical equilibrium is a dynamic state, both forward and reverse reactions continue, but because their performance is equal, there is no net change in the concentration of reaction components. What is a equilibrium position? Concentration of reactionary means and products in balance. What is the Le Chatelier principle? If stress is applied to a dynamic equilibrium system, the system changes in a way that relieves stress. List three stresses that can upset the balance of chemical reactions. - Changing concentration - Changing temperature- Changing pressure Energy that is available for work. List two characteristics that share spontaneous reactions. - Producing a large amount of product-free energy What is the law of disorder of the state? A natural trend for systems is to move towards an increase in disorder (entropy). How can you use entropy to determine whether a chemical reaction is more or less likely to be spontaneous? Reactions in which entropy increases as reactionary form products are likely to be favored. (Spontaneous). What two factors determine the spontaneity of the chemical reaction? - Size and direction enthalpy change-Entropy changes When Gibbs free energy changes for the process (-), what does this tell you about the process? If enthalpy increases and entropy increases as well, is the spontaneous reaction? Only if an unfavorable change in enthalpy is offset by a favorable change in entropy. If enthalpy decreases and entropy decreases as well, is the reaction spontaneous? Only if an unfavorable change in entropy is compensated by a favorable change of enthalpy. Which way will the equilibrium change if the pressure is increased? He will want to reduce the pressure and for this must increase the volume - that is, move to the side with fewer molecules. Which way would the equilibrium change if the pressure was reduced? He's going to want to increase the pressure and for that he has to reduce the amount -- then move aside with a large number of molecules. Which way will the balance change if there is an increase in the concentration of the product? It shifts in Side. Which way will the balance change if there is a decrease in the concentration of the product? It will move to the side of the products. Which way will the equilibrium change if you add heat? Heat can be considered a product/reaction depending on the chemical equation. If the heat is added, the equilibrium will shift towards the reactionary. Which way will the equilibrium change if the heat is removed? (Given that it is a product) it will move towards the product if heat is considered a product. It is a substance that increases the speed of reaction without being used during the reaction. They allow reactions to follow a lower energy path. What is Gibbs Free Energy? Energy? chapter 18 reaction rates and equilibrium answer key. chapter 18 reaction rates and equilibrium worksheet answer key. chapter 18 reaction rates and equilibrium guided reading answers. chapter 18 reaction rates and equilibrium answers pearson education answers. chapter 18 reaction rates and equilibrium packet answers. chapter 18 reaction rates and equilibrium pdf. chapter 18 reaction rates and equilibrium practice problems answers. chapter 18 reaction rates and equilibrium test a answers

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