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Heat of fusion of ice lab

molar heat of ice fusion = 6.03J/moleno depends on what the initial water temperature is, it is constant.the only thing that needs to be changed would be the time it took for the ice to melt molar ice fusion = 6.03J/mole, heat needed to melt all ice at the same temperature6.03kJ/mole x 1mole ice = 6.03kJ, for water to 6.03kJ energy, deltaT, water temperature, 14.4Cif one should use cold water to melt the ice, and this is done in a calorimeter where there is no heat transfer from the surrounding area, the ice would melt. the water must be warm enough to melt the ice on its own. Q lost water = Q obtained by ice melt, water temperature decreases, but the ice temperature remains at 0C until everything melts, then the temperature rises again as the total water reaches balance with surroundings.so, yes, the change in system temperature depends on the method used. cold water in the closed system did not melt the ice, while warm water in the closed system would melt the ice. Well for the data I wrote from the experiment I received: Weight calorimeter- 182 gMass calorimeter- 279 gMass warm water- 97 gMas calorimeter and all water / ice- 3 03 gMas added ice- 24 ginitial water temperature (maximum)- 50 degrees CFinal water temperature (minimum)- 26 degrees C They told me: 1. calculate the change in water temperature, delta T = Tmax-TminiI got +24 degrees C as my answer2. Calculate the heat(s) lost by water melting ice using equation Q=m(delta T)Cp. Cp for water = 4.18 J/g x C I received Q = -9,731.04 J3. Calculate the heat required to melt 1 g of ice (J/g)I have 405.46 J/g4. The accepted value for ice fusion heat is 334 J/g. Calculate your error percentage.% error = received value - experimental value / received value x 100 I have 40,400%, which I really need help on IS:5. What assumption did we make about the heat lost by water in the calorimeter compared to the heat obtained by melting ice? ANDExtension: Design an experiment to see if an ice cube taken from the freezer and immediately placed in a calorimeter needs the same amount of energy per gram to melt as an ice cube that has been out of the freezer for 10 minutes. Maybe for expansion can you please explain if the ice cube needs the same amount of energy per gram for melting? How and why, please? I really don't understand.}; Thanks for the help!

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