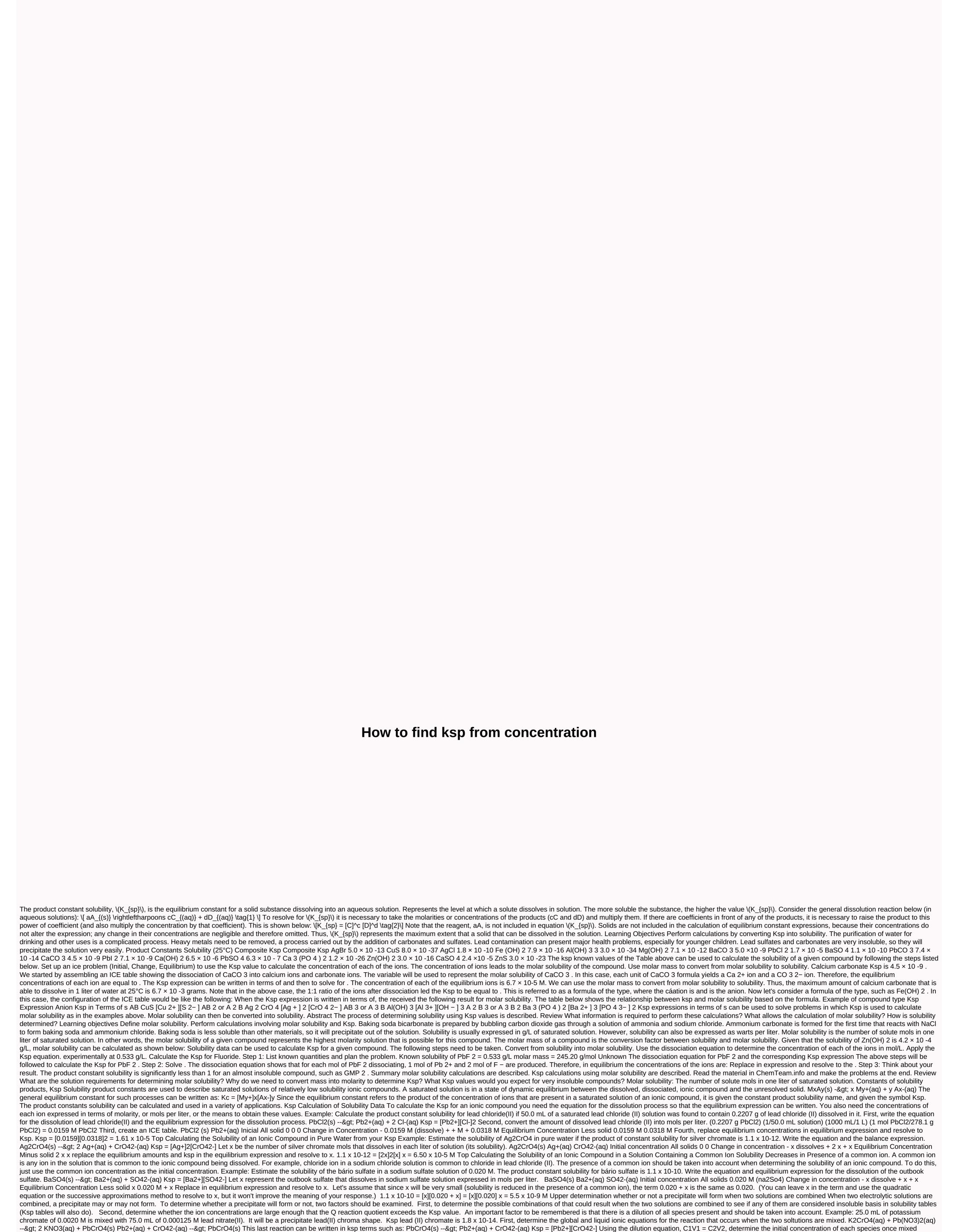
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(before any reaction occurs). (0.0020 M K2CrO4) (25.0 mL) = (C2)(100.0 mL) C2 to K2CrO4 = 0.00050 M Similar calculation for lead nitrate yields(II): C2 for Pb(NO3)2 = 0.0000938 M Using initial concentrations, calculate the reaction quotient Q, and compare with the value of the equilibrium constant,

Ksp. $Q = (0.0000938 \text{ M Pb2+})(0.00050 \text{ M CrO42-}) = 4.69 \times 10-8 Q$ is greater than ksp as soon as a lead precipitate chromate(II) is forsa. Top Top

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