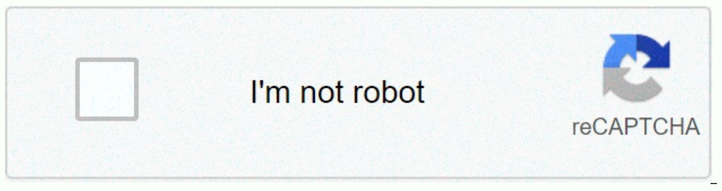

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Atomic structure chapter quiz answer key

We will thoroughly review each answer to the question to provide you with the correct answers. Did you find a mistake? Let us know by using the REPORT button at the bottom of the page. Atoms consist of three fundamental particles: protons, electrons and neutrons. The nucleus (center) of the atom contains protons (positively charged) and neutrons (without charge). The outermost regions of the atom are called electron shells and contain electrons (negatively charged). Atoms have different properties based on the arrangement and number of their fundamental particles. The hydrogen atom (H) contains only one proton, one electron and no neutrons. This can be determined by using the atomic number and the mass number of the element (see the concept of atomic numbers and mass numbers). Atomic Structure Quiz Question Your Answer: Correct Answer: You have {{SCORE_CORRECT}} of {{SCORE_TOTAL}} 1 Atomic Structure Quiz Answers 2 Terms & Answers Keys learn atomic structure Quiz Answers 1. All elements are composed of atoms of the compound of neutron molecule 2. What is the carbon mass (C)? 3. How many electrons are there in the carbon atom (C)? (Remember APE! Atomic number = # protons and # electrons) 4. How many electrons are in the carbon atom (C)? (Remember APE! Atomic number = # protons and # electrons) 5. How many neutrons are there in the sodium atom (Na)? (Remember M-A = N! Mass # - Atomic # = Neutrons) 6. What is the weight number of sodum (Na)? (Don't forget to round!) 7. Atoms of the same element shall always have the same number of ____ isotopes of electron proton neutrons 8. Dense, positively charged matter is located in the center of the atom. 9. __ an element equals the number of protons in an atom of that element by the atomic number of an atomic mass isotope of 10. ___ the atom is the sum of the protons and neutrons in the nucleus of that atom. atomic mass of the isotope atomic mass number 11. The number of protons in the element atom. atomic mass isotopes atomic number of protons 12. Where are the electrons in the atom? in the atomic number inside the nucleus outside the nucleus in electron zone 13. Where are the protons and neutrons in the atom? inside the nucleus in the electron zone outside the nucleus in atomic number 14. The element will always have the same number of ____. proton atoms by neutron isotopes 15. What is the atomic number of gold (Au)? 16. How many protons does gold (Au)? (Remember APE! Atomic number = # protons and #electrons) 17. Which atomic particles can you find inside the nucleus of the atom? Electron & Proton Proton & Neutron Neutron & Electron 18. Which atomic partied has a negative charge? Conditions & Answer Keys to learn the smallest partite element that maintains its chemical properties in the early 1800s came its own atomic model, which suggested that all matter is made up of individual particles called atoms that cannot be divided. 1. Elements consist of atoms 2. Atoms of the same element have the same mass (WRONG) 3. Compounds contain atoms of more than one element 4. in the compound, atoms always combine in the same way performed some experiments with electricity and matter. He found that electricity consists of negatively charged particles. negatively charged particles of positively charged particles of neutrally charged partimals have found that the core is positively charged because the protons are in the core. He did an experiment with gold foil. small particles that make up protons and NEUTRONS, each proton/neutron each has 3 up, down, top, bottom, strange, magic proton or neutron consists of ___ quarks equal to the sum of protons and neutrons in the atom, mass number – atomic number = # neutrons, we get the mass number by rounding the average atomic mass, the number of protons in the atom, the atom is identified as an element by the number of protons in the amu nucleus, based on the mass of one carbon atom that has 6 protons and 6 neutrons, 1 am = 1/12 the mass of the carbon atom amu is ___ the mass of the carbon atom protons + the mass of neutron atoms with the same number of protons but different number of neutrons Let us know if it was useful. That's the only way we can get better. Atoms are made of protons, neutrons and electrons. A periodic table organizes elements according to their atomic size and other properties. Properties.