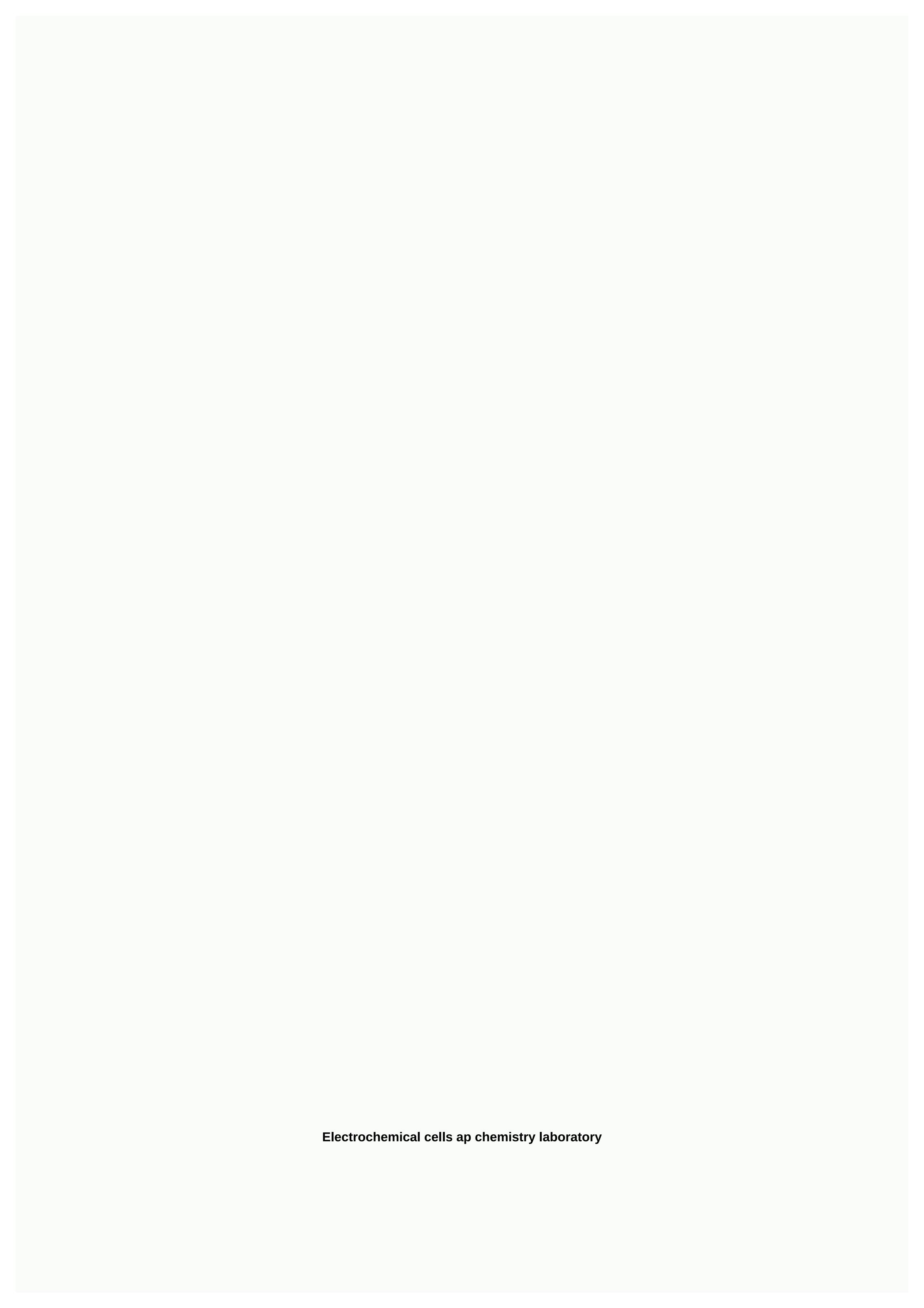
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The laboratory is carried out in three parts. In part 1 there is a table with a list of potentials for reducing metal ions. In part 2, the Nerst equation is used to measure cell tension. In part 3, the agcl solubility product constant is determined using the Nerst equation and voltic cells. To provide first aid, call your doctor and seek medical attention for further treatment, observation and support after first aid. Do not forget to wear an apron, glasses and gloves, if any. Copper (II) nitrate solution, odorlessSvitoktuto toxic when ingestedIt is immediately available to the skin, eyes and mucous membranesShow to induce, remove from the fresh air immediately; if breathing has stopped, give artificial breathing immediately for eye contact, immediately for eye contact, immediately wash with fresh water for at least 15 minutesFor external contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, give a large amount of water and call your doctor or poison control center immediatelyNext, non-combustible solutionFor spills, contain it with sand or absorbent material and put it in an airtight bag or containerIron (III), 1.0 MBrown-yellow liquid, odorlessMay be an irritant of the skin and tissueAoidal contact with the skin, eyes and mucous membranesHow to induce, remove in fresh air immediately; if breathing has stopped, immediately give artificial breathing for eye contact, immediately wash with fresh water for at least 15 minutesFor external contact, err errand the infected area with fresh water for at least 15 minutesFor internal contact, give a large amount of water and call your doctor or poison control center immediatelyNext, non-combustible solutionFor spills, container, . Mg(NO3)2, 1.0 MClear, colorless solution, odor-free irritant tissueAoid contusion with all body tissuesHow to inhale. remove from fresh air at once; if breathing has stopped, immediately give artificial breathingfor contact, rinse one or two cups of water or milk, and cause vomiting. Call your doctor or poison control center immediatelyNegortal, non-combustible solutionFor spills, pour it with sand or absorbent material and put it in an airtight bag or containerLead (II) nitrate solution, Pb(NO3)2, 1.0 M When inxitating, remove it to fresh air immediately; if breathing has stopped, immediately give artificial breathingfor eye contact, immediately rinse with fresh water for at least 15 minutes For internal contact, give a large amount of water and or poison control center immediately rinse with fresh water for at least 15 minutes For internal contact, give a large amount of water and or poison control center immediately For spills, contain it with sand or absorbent material and put in an airtight bag or containerSouth nitrate solution, AgNO3, 1.0 MClear, colorless your solution, odorlessly toxic when ingested irritating contusion fabric with all tissues of the bodyHow to inflate, remove if breathing has stopped, immediately give artificial breathingfor contact with the eyes, immediately erne errand fresh water for at least 15 minutesFor external contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine the infected area with fresh water for at least 15 minutesFor internal contact, ermine for at least 15 minutesFor internal co put it in an airtight container, Zn(NO3)2, 1.0 MClear, colorless solution, odorless World toxic when corrosive in the body tissues of the body tis water for at least 15 minutes For external contact, rinse the infected area with fresh water for at least 15 minutes For internal contact, rinse your mouth, give one or two cups of water or milk for dilution. DO NOT cause vomiting. Call your doctor or poison control center immediately Neconcombial solution For spills, contain it with sand or absorbent material and settle in an airtight bag or container Sodium chloride solution, NaCl, 1.0 MColorless aqueous solution, Odorless Hazard in high concentrationsNew for human consumption or use the topicHow to inflate, remove to fresh air immediately; if breathing has stopped, immediately give artificial breathing for eye contact, immediately wash with fresh water for at least 15 minutesFor external contact, erine your mouth and give plenty of water for dilutionControl liquidFor spills, hold it with sand or absorbent material and put it in an airtight bag or container Toussia nitrate solution. , KNO3, 1.0 MColo colorless your solution, odorless Toxic levels have not yet been studied When indented, remove it from fresh air immediately; if breathing has stopped, immediately give artificial breathing for eye contact, immediately rinse with fresh water for at least 15 minutesFor external contact, rinse the infected area with fresh water for at least 15 minutesFor internal contact, rinse your mouth and give plenty of water for dilutionNeconcommunicable liquidFor spills, hold it with sand or absorbent substance and put in an airtight bag or container Short foilDo not absorb in any wayNot should play with metal metal on other objects or organismsUsing common sense when using metal at all timesMain fabric; in case of err 2000s thum area and put the bandage, if it is appropriateIron nailNew to swallow in no wayNe play with metal Do not do horseback riding with metallt is worth clamping the metal on other objects or organismsUsing common sense when using metal at all timesMay cut fabric; when bleeding, err on the site and put a bandage, if appropriateUse not to absorb in any way, so as not to play with metal, do not do horseback riding with metallt is worth clamping the metal on other objects or organismsUsing common sense when using metal at all timesMait cutting fabric; when bleeding, err on the site and put the bandage if necessaryMagnion not to swallow, so as not to flysage with metal, do not do horseback riding with metal, do not clamp the metal on other objects or organismsUsing common sense when using metal at all timesMain fabric; in case of bleeding, err on the area and put the bandage, if it is appropriateSouth foil, do not swallow in any way Do not play with metal Do not play a riding game with metal timesMain fabric; in case of bleeding, err on the area and put the bandage, if it is appropriateSouth foil, do not swallow in any way Do not play with metal Do not play a riding game with metal timesMain fabric; in case of bleeding, err on the area and put the bandage, if it is appropriateSouth foil, do not swallow in any way Do not play with metal Do not play a riding game with metal timesMain fabric; in case of bleeding, err on the area and put the bandage, if it is appropriateSouth foil, do not swallow in any way Do not play with metal Do not play with metal timesMain fabric; in case of bleeding, err on the area and put the bandage, if it is appropriateSouth foil, do not swallow in any way Do not play with metal Do not play with common sense when using metal at any timeMay the cut fabric; in case of bleeding, rinse the area and put a bandage, if it is appropriateThe arrangement of stripteaseDotra do not absorb in any way Do not play with metal, do not do horseback riding with metalNe should you squeeze the metal on other objects or organismsUse common sense when using metal at all timesMay the cut fabric; in case of bleeding, ermine the site and put the bandage if necessaryCoper (II) of nitrates, Pb(NO3)2, 1.0 M, 2 mlMagnet nitrate solution, Mg(NO3)2, 1.0 M, 2 mlMagnet nitrate solution, Mg(NO3)2, 1.0 M, 2 mlMagnet nitrate solution, Mg(NO3)2, 1.0 M, 2 mlMagnet nitrates, Pb(NO3)2, 1.0 M, 2 mlMagnet ni mlSylver nitrate solution, AgNO3, 1.0 M, 2 mLZinc nitrate solution, Zn(NO3)2, 1.0 M, chloride solution 2 mlSudium, NaCl, 1.0 M, nails 10 mlIron, feLead foil, PbMagnesium tape, Foil MgSilver, AgZing strip, ZnReaction plate, 24-cylinder, 10 mlWires and alligator clipsThe same baths, 12 x 150 mm, 3Steel wool or aussage paperBeral pipe type, graded 10Filter paperVoltmeterBeaker, 50ml Materials have been installed. With the help of a pipette, 2 ml of zinc nitrate was dropped into the well, and 2 ml of copper (II) nitrates were put in a well next to zinc nitrate. With the help of sandpaper, strips of zinc and copper were polished. Small pieces were cut that could fit in the well. A zinc piece of metal was attached to a positive terminal, while copper was attached to a negative voltmeter terminal, while doing this, the salt bridge was prepared by immersing a rectangular piece of filter paper in a solution of potassium nitrate. Then zinc metal was immersed in copper (II) solution. The tension was then received by looking voltmeter display. This procedure was repeated four more times by disabling copper with other metals: silver, iron, magnesium and lead, which were immersed in silver nitrate and lead (II) nitrate solutions, respectively. Later, metal nitrates were placed in the following order on wells: New salt bridges made of potassium nitrate were placed vertically through the upper wells to the wells directly to them. Then the metals were connected to a voltmeter and immersed in wells. First there was copper and magnesium, in the far left of the well a plastic block, and then iron and copper, etc. until copper and silver steam were reached. Tensions were fixed. For part 2, copper solution (II) 0.0010 M was manufactured by diluting 1.0 M of solution. 2 drops of 1.0 M each containing 18 ml of distilled water were added to the test sampling. It was thoroughly mixed, and the process was repeated two more times to purchase a solution of 0.0010 M. Part of this solution was put in a well, and a piece of copper metal was immersed in a well, and a piece of copper metal was immersed in a well, and the result was fixed. Ten millimeters of one sodium molar chloride was poured into a fifty millimeter beak. Then one drop of silver nitrate was added, and the stirrer was mixed. This led to the formation of some AgCl sediment. The solution was poured into the well, and then zinc nitrates, and silver, from which a small piece of metal was immersed in a previously prepared solution. The results were recorded, Part 1Voltage of each half of the cell vs. Zinc ElectrodeVoltageAnodeCathodeZn vs. Fe0.31 VZnFeZn vs. Mg-0.57 VMgZnZn vs. Pb-0.28 VZnPbEdited and measured cellular potential Anode Cathode Equation for cellular reaction Ed potential from experimental dataMeasured Potential MgCu2+(aq) + Mg(s)  $\rightarrow$  Mg2+(aq) + 2Efe(s)1.11 V-0.42 VFeAg 3Ag+(aq) + Fe(s)  $\rightarrow$  Fe3+(aq) + 3Ag(3Ag)s) 1.57 V-0.89 VMgPb Pb2+(aq) + Mg(s)  $\rightarrow$  Mg2+(aq) + Mg(s)  $\rightarrow$  Mg2+(aq) + 2Efe(s)1.11 V-0.42 VFeAg 3Ag+(aq) + Fe(s)  $\rightarrow$  Fe3+(aq) + 3Ag(3Ag)s) 1.57 V-0.89 VMgPb Pb2+(aq) + Mg(s)  $\rightarrow$  Mg2+(aq) + Mg2+(aq)  $\rightarrow$  Mg2+(aq) + Mg2+(aq)  $\rightarrow$  Mg2+(  $Mg2+(ag) + Pb(s)2.24 V0.22 VPbCu2+(ag) + Pb(s) \rightarrow Pb2+(ag) + Cu(s) 0.47 V0.48 VCuAg 2Ag+(ag) + Cu(s) \rightarrow Cu2+(ag) + Cu(s) \rightarrow Cu2+(ag) + Cu(s) 0.47 V0.48 VCuAg 2Ag+(ag) 0.47$ For part one, the reduction equation for each metal ion, from the lowest measured potential of the electrode with the permissible potential of an electrode using hydrogen as a standard. The difference between them is shown. Reduction Equilibrium Electrode Potential using Zinc as the Standard, E°ZnAccepted Electrode using Hydrogen as Standard, E°ZnAccepted E 2e— → Pb(s) -0.28 V -0.13 V -0.15 V Mg2+ + 2e— → Mg(s) -0.57 V -2.37 V 1.80 V Zn2+ + 2e— → Zn(s) 0 V -0.76 V 0.76 V . For Part 2, the balance equation for the reaction occurring in the cell is written. Then the Nerst equation is used to calculate the expected voltage to compare with the measured voltage. Equation: Cu2+(aq) + Zn(s)  $\rightarrow$  Cu(s) + Zn2+(aq)Expected voltage: E = E°cell - (0.0592/n)logQE°cell = Ered(cathode) - Ered(Ered) (Ered) anode) E°cell = 1.10 vn = 2Q = [Zn2+(aq)]/[Cu2+(aq)] = (1.0 M)/(0.0010 M) = 1000logQ = log(1000) = 3E = 1.10 in - (0.0592/2)•3E = 1.1 089E = 1.0 v Compare with measured voltage: Measured: 0.82 vPecpected: 1.0 vDyencence: 0.18 v For part 3, the imon equation is used to calculate the concentration of an Ag+ ion. This is done by finding the AgCl solubility product using the equation. The calculated solubility product is then compared to the actual value. AgCl concentration was also detected. Equation: 2Ag2+(ag) + 2Ag(s)Calculated Ksp:On Equilibrium, value E is zero.0 = E°cell - (0.0592/n)logO = E°cell = -0.72 v(0.0592/2)logO = 0.72logO = 24.32O = 2.4 • 1024Q = KspKsp= 2.4 • 1024Computer to actual value: Calculated: 2.4 • 1024Active: 1.77 • 10–10Indance: 2.39 • 1024AgCl M:Ksp = [Products]/[Reagents]2.4 • 1024]1/2[Ag2+] = 6.5 • 10-13 M The reaction potential of various electrodes and reactions was calculated using cells. They were compared with the real reaction potentials. The expected value of E in the reaction between copper and zinc was set to 1.0 c. The difference between the measured in the laboratory and the expected value is 0.18 v. This is by 2.39 • 1024 of the expected value, 1.77 • 10–10. AgCl concentration was recognized at 6.5 • 10-13 M. The difference between the measured values and the actual values of the potentials for reducing metals and reactions was very different from the expected values. This may be due to the fact that the metal is not sufficiently immersed in wells – when performing the laboratory, small changes in the position of the metal inside the well led to the fact that the voltage Many. This applies to all measured reaction potentials. Since they were not accurate, the accuracy of other calculations dependent on reaction potentials also became inaccurate, as evidenced by the large difference between the calculated solubility product with the actual value. MediaFlinn Scientific, Inc. - Technical Letter on Material Safety (MSDS). Flinn Scientific, Inc. - Materials Security Data (MSDS). Flinn Scientific, Inc. PDF. II) NitrateSol.pdf, Iron(III) NitrateFlinn Scientific, Inc. - Materials Security Data (MSDS). Flinn Scientific, Inc.PDF. 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