

☐

I'm not robot


reCAPTCHA

Continue

Single and double replacement reactions worksheet answers

Back to Double Replacement Write correct formulas for the products in these double replacement reactions. 1) $\text{Ca(OH)}_2 + \text{H}_3\text{PO}_4 \rightarrow$; $\text{Ca}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$ 2) $\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow$; $\text{KCl} + \text{BaCO}_3$ 3) $\text{Cd}_3(\text{PO}_4)_2 + (\text{NH}_4)_2\text{S} \rightarrow$; $\text{CdS} + (\text{NH}_4)_3\text{PO}_4$ 4) $\text{Co(OH)}_3 + \text{HNO}_3 \rightarrow$; $\text{Co(NO}_3)_3 + \text{H}_2\text{O}$ 5) $\text{AgNO}_3 + \text{KCl} \rightarrow$; $\text{AgCl} + \text{KNO}_3$ 6) $\text{Na}_2\text{CO}_3 + \text{H}_2\text{SO}_4 \rightarrow$; $\text{Na}_2\text{SO}_4 + \text{CO}_2 + \text{H}_2\text{O}$ 7) $\text{Al(OH)}_3 + \text{HC}_2\text{H}_3\text{O}_2 \rightarrow$; $\text{Al(C}_2\text{H}_3\text{O}_2)_3 + \text{H}_2\text{O}$ 8) $\text{Al}_2(\text{SO}_4)_3 + \text{Ca}_3(\text{PO}_4)_2 \rightarrow$; $\text{AlPO}_4 + \text{CaSO}_4$ 9) $\text{Cr}_2(\text{SO}_3)_3 + \text{H}_2\text{SO}_4 \rightarrow$; $\text{Cr}_2(\text{SO}_4)_3 + \text{SO}_2 + \text{H}_2\text{O}$ 10) $\text{AgC}_2\text{H}_3\text{O}_2 + \text{K}_2\text{CrO}_4 \rightarrow$; $\text{Ag}_2\text{CrO}_4 + \text{KC}_2\text{H}_3\text{O}_2$ Here are more examples of the decomposition into gas and water: $\text{KOH} + (\text{NH}_4)_2\text{SO}_4 \rightarrow$; $(\text{NH}_4)_3\text{PO}_4 + \text{Mg(OH)}_2 \rightarrow$; $(\text{NH}_4)_2\text{SO}_4 + \text{Ca(OH)}_2 \rightarrow$; $\text{MnO}_2 + \text{HCl} \rightarrow$; $\text{MnCl}_2 + \text{Cl}_2 + \text{H}_2\text{O}$ Actually, this last one is a bit of an oddball at this beginning level. The ChemTeam wouldn't ask it for a test, but the other three would be fair game. Back to Single Replacement Write correct formulas for the products in these only replacement reactions. Unbalanced response 1) $\text{Al} + \text{Pb(NO}_3)_2 \rightarrow$; $\text{Pb} + \text{Al(NO}_3)_3$ 2) $\text{Cl}_2 + \text{NaI} \rightarrow$; $\text{I}_2 + \text{NaCl}$ 3) $\text{Fe} + \text{AgC}_2\text{H}_3\text{O}_2 \rightarrow$; $\text{Ag} + \text{Fe(C}_2\text{H}_3\text{O}_2)_2$ 2 If Fe(III) is used, then $\text{Fe(C}_2\text{H}_3\text{O}_2)_3$ would result. 4) $\text{Al} + \text{CuCl}_2 \rightarrow$; $\text{Cu} + \text{AlCl}_3$ 5) $\text{Br}_2 + \text{CaI}_2 \rightarrow$; $\text{I}_2 + \text{CaBr}_2$ 6) $\text{Al} + \text{HCl} \rightarrow$; $\text{AlCl}_3 + \text{H}_2$ 7) $\text{Mg} + \text{HCl} \rightarrow$; $\text{MgCl}_2 + \text{H}_2$ 8) $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow$; $\text{ZnSO}_4 + \text{H}_2$ 9) $\text{Fe} + \text{CuSO}_4 \rightarrow$; $\text{Cu} + \text{FeSO}_4$ If Fe(III) is used, then $\text{Fe}_2(\text{SO}_4)_3$ would result. 10) $\text{Cl}_2 + \text{MgI}_2 \rightarrow$; $\text{I}_2 + \text{MgCl}_2$ Here are examples of the +1 and +2 cations reacting with water: $\text{Na} + \text{H}_2\text{O} \rightarrow$; $\text{NaOH} + \text{H}_2$ $\text{Li} + \text{H}_2\text{O} \rightarrow$; $\text{LiOH} + \text{H}_2$ $\text{Cs} + \text{H}_2\text{O} \rightarrow$; $\text{CsOH} + \text{H}_2$ $\text{Mg} + \text{H}_2\text{O} \rightarrow$; $\text{Mg(OH)}_2 + \text{H}_2$ $\text{Ca} + \text{H}_2\text{O} \rightarrow$; $\text{Ca(OH)}_2 + \text{H}_2$ $\text{Ba} + \text{H}_2\text{O} \rightarrow$; $\text{Ba(OH)}_2 + \text{H}_2 + \text{H}_2$