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Напечатано: PRINTED В SPAIN - PRINTED В SPAIN Эта книга была напечатана с органической бумагой и inkwww.FreeLibros.orgResumen содержания Первая часть Введение и справочная информация 2 Глава 1 Природа экологии 4 Глава 2 Адаптация и эволюция 17 Часть вторая Физическая среда 44 Глава 3 Климат 46 Глава 4 Водная среда 68 Глава 5 Наземная среда 87 Третья часть Организм и окружающая среда 106 Глава 6 Адаптация растений к окружающей среде 108 Глава 7 Адаптация животных к окружающей среде 139 Глава 8 Образец жизненного цикла 172 Часть 4 Население 194 Глава 9 Свойства населения 196 Глава 10 Рост численности населения 215 Глава 11 Внутриспецифическое регулирование народонаселения 235 Глава 12 Целевой 255 Пятый Взаимодействие между видами 270 Глава 13 Межспецифическая конкуренция 272 Глава 14 Предтеча 299 Глава 15 Тунелство и взаимность 327 part of his Community Ecology 348 Chapter 16 Community Structure 350 Chapter 17 Factors, Affecting Community Structure 370 Chapter 18 Dynamics Community 391 Chapter 19 Landscape Ecology 416 Seventh Part Ecosystem Ecology 442 Chapter 20 Ecosystem Energy 444 Chapter 21 Nutritional Breakdown and Circulation 470 Chapter 22 Biogeochemical Cycles 496 Eighth Part BiogeographicAl ecology 518 Chapter 23 Terrestrial Ecosystems 520 Chapter 24 Aquatic Ecosystems 546 Chapter 25 Ground Water Transitions 569 Chapter 26 Large-scale Biodiversity Models 585 Ninth Part Ecology 598 Chapter 27 Population Growth, Resource Use and Sustainability 600 Chapter 28 Habitat Loss, Biodiversity and Conservation 630 Chapter 29 Global Climate Change 653www.FreeLibros.org-Summarycontainment vww.FreeLibros.orgContenidoPreface xviii 2.4 Genetic variation is an important ingredient for natural selection 28First batch Introduction and 2.5 Evolution is a modification in laantecedents 2 gene frequency 29Caption 1 Nature Ecology 4.2.6 The concept of species based on genetic isolation 30 1.1 Organisms interact with the environment in the context of ecosystem 4 2.7 The process of species includes the development of reproductive isolation 32 1.2 Ecosystem components form a hierarchy 5 2.8 Geographical differences in the form provides information on the process 1.3 Ecology has complex roots 5 species 34 1.4 Ecology has links 2.9 Adaptation reflects the obligations and disciplines of 7 constrictions 35 Profile researchers Beren W. 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that falls to the ground in each one is different: in colder waters, there are streams on the surface and filters into the ground, and thus get more oxygen, such as trout and solutions. Surface waters such as the creeks of the Icelandic American small mouth; and in warmer waters, rivers, collect more solutions of substances and above them are species that require less oxygen that flow. Water of most rivers and lakes with such folds and som (see figure 24.13). dissolved minerals. The relative concentrations of minerals in these waters reflect the substrates on which they flow. For example, water flowing over areas of 4.5 Water functions as a solvent, which largely underlies the basis of limestone (consisting mainly of calcium carbonate; CaCO3) have a high con-if a tablespoon of sugar is removed in a glass with calcium (Ca2) and bicarbonate (HCO3) water, sugar dissolves, and homome-nea mixture is formed. Compared to fresh water, a pre-substance of the oceans, a liquid that is homogeneous mixture of two or more is called dissolution. The solvent agent sits much higher concentration soluble. Deve dissolution solvent or solvent, and substance made. oceans function as large yet. The one who dissolves is called soluble. Dissolution, the flow of fresh water into the oceans of which is constantly increasing - is water, called acquiescence. soluble water content, as clean water evaporates from the surface and passes into the atmosphere. Water is an extraordinary solvent that has a concentration of solvents, however, it cannot continue the ability to dissolve more substances than any other lipid growth indefinitely. When the concentration dequido. This is Extraordinary function as some elements reaches the limit set by ladisolvente makes the water biological substance the maximum solubility compounds they form (critical grams. This is a liquid in which molecules per liter), excesses are deposited and deposited as sedatives and waste can dissolve and cops. Calcium, for example, easily forms a carb transport, helps regulate temperature and stores Nato calcium (CaCO3) in ocean waters. Chemical balance in living cells. The maximum solubility of calcium carbonate, however, is only 0.014 grams per liter of water, the concentrate-ability of water to function as a solvent due at the beginning of the history of the ocean, mainly to the link covered by section 4.2. As a result, calcium ions are deposited by the continuous result of the asymmetrical bond of atom H with O te atoms from the solution and deposited at the bottom of the ocean (see figure 4.3), one side of each water molecule has both limestone deposits, constant positive load, and the other side has a constant payload; This situation is called dipole In contrast, the solubility of sodium chloride is very (fields with opposite loads). Because the opposite loads are high (360 grams per liter). In fact, these two elements, in addition to attracting each other strongly, mole-sodium and chlorine, form approximately 86 percent of those from the water also attract other molecules with load. sea salt. They, along with other important elements such as sulfur, magnesium, potassium and lime- Compounds consisting of atoms of okrium groups, whose relative proportions differ little, are electrically charged atoms called ion. 99 per cent sea salt (table 4.1). Sodium chloride (table salt) definition, for example, is one of the most common elements, chlorine used as an index formed by positively charged sodium ions (NAH) salinity. Salinity is expressed in negatively charged carbonated blocks of chlorine (Cl-), located in for-functional (USF) (represented as a), measures such as the crystalline network. When salt is poured into the water, grams of chlorine per kilogram of water. The salinity of maratractions between negative (oxygen atom) and open loads is quite constant with 350 on average. Imposives (hydrogen atoms) in the water change molecule, freshwater salinity varies between 0.065 (see figure 4.3), as well as sodium and chlorine atoms and 0.30. However, over time more than the forces held together by chrysiological salt shale (ion bonds). As a result, the ocean crystals have increased and haciendolo de easily dissolves in the ions that connect them to the water, i.e. dissolve. Water solvents are the cause of most minerals (elements and compounds inor- 4.6 Oxygen is dissipated from tannins) found in the aquatic environment. Atmosphere to surface waters When water condenses and forms clouds, it is almost pure, salty part of some dissolved atmospheric gases. When water is not limited to dissolving the surface in the form of precipitation, it acquires solids. The surface of the reservoir determines the temperature of unwww.FreeLibros.org-76 SecondPart Physical EnvironmentTable 4.1 Composition of sea salt 35 units the speed at which the gases are dissipated by water. Functional salt gas scattering (USF) occurs about 10,000 times more in water than in the air. In addition to the diffusion process, elements of g/kg Milli-Milli-Oxygen, absorbed by surface waters, mix moles/kg equivalents/kg with deeper water thanks to internal currents and aCationes 10,752 467.56 467.56 turbulence. In shallow water worksOidium 0.395 10.10 10.10 very quickly and in wind-induced deuces, oxygenPotassium can reach and maintain saturation levels and even oversaturation by increasing the absorption area in laMagnesium 2,295 53.25 106.50 contact between the surface air and water. Oxygen heated 0.416 10.38 20.76 loses in water when the temperature rises, decreases-strontium 0.008 0.09 0.18 th solubility and consumption of aquatic life. In summer, oxygen, as well as temperature (see - 605.10, section 4.4), can be stratified into lakes and lagoons. The largest amount of oxygen is usually given near the surface of the information, where there is an exchange between waterCloro 19,345 545.59 545.59 and the atmosphere, stimulated by the removal of the action TheBrome 0.066 0.83 0.83 wind (Figure 4.7). In addition to its entry into the water at 0.0013 0.07 0.07 average diffusion of the atmosphere, Oxygen is also a product of photosynthesis, severely limited in sodium 2701 28.12 56.23 surface water due to limitations in the availability of bicarbonate 0.145 2.38 - light (see Figure 4.5 and Chapter 6). The amount of oxy-boric acid 0.027 0.44 - geno decreases with depth due to the demand for oxygen decomposing organisms living on 602.72 bottom of sediments (chapter 21). During the prima faith and renewal of the odonya water, when water circulates across the lake, oxygen is replenished in the deep, bordering on the atmosphere. The gases are exchanged through winter, reducing oxygen in water that is not the master of this limit during the diffusion process. Diffusion is frozen mild, as the demand for oxygen for the total loss of molecules to move out of the body is reduced by cold, and the capacity of high concentrations of oxygen-region at low concentrations is not higher at lower temperatures. There is no ice (see quantitative ecology 4.1: Diffusia and Osmos). However, the decrease in oxygen can be serious due to the diffusion process produces global transmission, the lack of diffusion from the atmosphere to the superfi-two important gas from a metabolic point of view, ciales, oxygen and carbon dioxide, from the atmosphere (con- As in lagoons and lakes, oxygen does not emphasize above) to surface waters (concent-distributes evenly in the depths of the lower oceans) the environment. years (Figure 4.8). The typical oxygen profile of the Ocean Oxygen dissipated from the atmosphere into the water shows the maximum amount in the upper 10-20 m surface. The diffusion rate is controlled by where photosynthetic activity and diffusion from oxygen lasolyuati in water and the slope of the gra atmosphere usually lead to saturation. In diffusion eldient (difference in concentration of elevated depth, dismi-air oxygen content and surface water where diffusion occurs). He sniffs it. In open ocean waters, concentrations of waterproof gas are a function of temprat-reach a minimum value, somewhere between 500 yra, pressure and salinity. The value of oxygen saturation is 1000 m, an area called conteni-core area in cold water than in hot water, because minimal oxygen is solubili-do. Unlike lakes and laga-papa (the ability to remain as a solution) gas in us, where seasonal decomposition of thermoline and water decreases when the temperature rises. Without the mixture that leads to between the surface and deep water, solubility increases with increased lead to dynamic temperature gradient and atmospheric depression and decreases with increasing oxygen content, limited depth of super-nity, which has little value in fresh water. Mixing fiss in deep oceans supports gra-in surface waters, the diffusion process continues and the vertical availability of tooth oxygen all year round, oxygen dissipates from surface to water that the presence of oxygen in lower water (lowest concentration). The largest densi-moving water is quite different. Agitation and viscosity of water associated with air serve to limit and constantly swirl watercourses comwww.FreeLibros.org-Chapter4 Ambient 77When ecology is quantitatively 4.1 diffusion and osmos diffusion is a passive global movement of particles such as ha-ses such as oxygen (atoms, ions or molecules) from high con-area and carbon dioxide to lower concentrations. In others, they dissipate from the atmosphere into the surface layer of words, any substance dissipates, and therefore reduces water in aquatic ecosystems. Once they're in laye, their concentration gradient. One substance makes this surface layer, these gases continue to spread to a two-way that reaches balance (Figure 1). This one is reached downstream. We can use this equation when the concentration is equal at all points to calculate the diffusion rate of the carbon dioxide system. between the surface layer of water and the depth of 10 cm as follows: Fka's law describes the flow (sus-tance movement): Flow or movement of the diffusion factor. The diffusion rate of CO2 CO2 concentration at one point a (mol/cm2/s) (mol/cm3) (mol/cm3) describes how quickly it is transferred over time to another. They are distributed by substance J (1.6 x 5 x 5 x 10-5 x 2.9 x 10-5)/10 x 4.32 x-11 units of mole per unit area and through Wednesday. Units on time. For example: unit time: cm2/s. mol/cm2/s. Distance diffusion ratio (see) for CO2 (cm2/s) J x D ΔΔz Concentration Ratio. The distance at which he was (22 - focused in addition to being An important process for mo- 1 viton substances in and between the atmosphere and morning- 3 water wells (such Δ as the movement of substances in C2), therefore, even at saturation levels, it is scarce and problematic. A slight decrease in the value of genoenoenza in water reaches a maximum of 0.01 liters per liter in deep wells or in contaminated water. (1%) fresh water at 00c. As a result, the concentration of oxygen in the aquatic environment, however, even in ideal conditions, solubi- in general, limits the breathing and metabolic activity of gases in the water no more. For example, oxy-geno is almost never confined to the Earth's environment. The aquatic environment, however, provides oxygen, (see section 6.1).www.FreeLibros.org-78 Secondpart Physical Environment occurs in biological processes. Most of the transport pattern 2a through cell membranes is done by blurring (see chapters 6 and 7). When osmotic matter is involved in maintaining a balance concentrated more on one side of the membrane than on the other, there is water in all living beings, from the simplest to the tendency for the substance to spread through one cell of vertebrates. Osmosis occurs with onemembrane, reducing its role as a concentration gradient is especially important in organisms (provided that the membrane is permeable to this substance). Fresh water and marine, whose concentrations Of Caco, membranes permeable internally soluble differ from water to subselective, and therefore affect the dife-renal rate of dife-leased substances. water diffusion through the disatu shift membrane- 1. What is different about the influence of paravameable mosis, is a special case of transporting passi-fish, inhabited in freshwater and paravo environments called mosses. Suppose we lock salt water? Suppose a desert concentration, such as salt (sodium chloride) in the soluble concentration (salt) of fish tissues, is located between delillate (and low-concentration water) in freshwater duct and salt water, sealed with a semi-permeable membrane and immersed in a glass of sediment with distilled water (Figu- 2. The membrane is permeable for water, but not for salt. how would the rate of blur change the volume of liquid inside the vessel increases and increases the injury (in the waters at 10 cm deep) you duplicate it through the tube, while the water moves through the mem-tion concentration of carbon dioxide in the elbran to the dissolution process called surface water? The water continues to move through the membra-na until the osmotic pressure of the solution (which decreases- you while the clean water dilutes the solution more) equals the physical pressure down exerted by the water column in Glass. The tendency to dissolve to carry water molecules from areas with high concentrations in an area with low concentration is called osmothical. The osmotic potential of the solution depends on its concentration. The higher the soluble conken tariff, the lower its osmotic potential and the greater its propensity to increase water. Moss is important for biological processes. Transporting water and other molecules through biological membranes is important for many pro-cesses in living organisms. Usmos and the potential4.7 Acidity has a large surface area, carbon dioxide reacts with water and influences the aquatic environment produces carbon acid (H2CO3). Carbon dioxide soot is slightly different from CO2 and H2O: H2CO3of oxygen, as it reacts chemically with water. Carbon acid, in turn, divides and forms ion water capable of absorbing hydrogen and bicarbonate ion, carbon dioxide and, therefore, the latter abounds in both fresh water and salt water. After spreading in H2CO3 ;: HCO3 and Hwww.FreeLibros.org-Chapter4 Environment 79Verenity (m) 0 Ice Summer surface ice reduces figure 4.7 Stratification 1 Winter oxygen diffusion from oxygen to Mirror Lake, N2 Atmospheric Renewal toward Hampshire Water, Winter, Summer and 3 Surface Autumn. late fall. Later call update 4 leads to 5 15 20 0 2 4 6 8 10 12 Oxygen Descent reflects the constant temperature and 60c) Oxygen (PPM) demand and oxygen distribution consumption of 7 homogeneous decomposing organisms in Pool 8 that inhabit the lower region. Lake. In summer, there is a 9 pronounced bundle of 10 temperature and oxygen. 110 5 oxygen is drastically reduced in the thermocline and does not exist at the bottom, due to the consumption of decomposing organisms in sediments. In winter, oxygen also stratified, but in deep water it exists in low concentration. (Adapted from Likens 1995.) 0 Carbon acid zone and bicarbonate produce more CO2, up to 1000 minimum new balance. Oxygen Chemical reactions are shown to result in inProfundity (m) production and absorption of free hydrogen ions (H+). The abundance of hydrogen ions in dissolution 2000 way of measuring acidity The higher the amount of H ions, the more acidic dissolution. Alkaline- 3000 Nas solutions are those that have a large amount of OH (hydro-xylos) and several H ions. Measuring the acidity and alkalinity of 4000 pH, which is calculated as a negative logarite (in 10) concentrations of hydrogen ions in the dis- 5000 0 lute. In clean water, a small fraction of the molecules are 1234 5 6 breaks down into ions: H2O: H 'OH, and the ratio of CM3 O2 per liter of H ions in relation to OH ions is 1:1. From both Figures 4.8 vertical oxygen profile at depth appears at concentration of 10-7 molecules per liter in the tropical Atlantic Ocean. The oxygen content in the neutral solution has a pH 7 -journal (10-7) x 7 .waters descends to a depth known as the area of dissolution is no longer neutral when the ion increases and contains minimal oxygen. Increased oxygen is thought to reduce others. Typically, logarithmdebajo of this area is the result of the arrival of cold, negative hydrogen water to classify lampaic solutions into oxygen coming from the immersion of acidic or major polar waters. Thus, an increase in hydroiron (see section 4.8). Geno 10-6 molecules per liter means a reduction of OH-ions to 10-8 molecules per liter, and the pH of the solution bicarbonate can, at the same time, disperse and form another 6. Negative logarithic scale ranges from 1 to 14. One is a hydrogen rhat and the other is carbonate ion. More than 7 denotes alkaline solution (more OH-concent), and pH less than 7, acidic dissolution. HCO3 ;: HK CO32 - Although clean water is pH neutral because the decomposition of water molecules produces the same carbon dioxide-carbon acid-bicarbo-number-number of ions H and OH-ions, the presence of CO2 enate is a complex chemical system that tends to permane-water alters this relationship.cer in balance. (Note that the arrows in the equations of the above reactions flow into production and are presented in both directions). As a result, the absorption of free hydrogen ions (ICD). Since the last is removed from the water, the balance is disturbed, and the equation-abundance of hydrogen ions in the solution is a measurement that has been presented to go left; The acidity yes, the dynamics of carbon dioxide-system-www.FreeLibros.org-80 SecondPart Physical Environment100 Figure 4.9 Theoretical CO percentages in each of the three co2 free 2A total CO2 percentage 50 HCO3 CO32 - water condition relative to pH. At low pH (acidic conditions) most of the CO is in 2 of its free form. In intermediate values (neutral conditions) bicarbonate prevails, while in alkaline conditions most CO is in the form of carbonate ions. 2 0 4 5 6 7 8 9 10 11 12 pHcIdo carbonic has a direct impact on pH embargo, as pH goes down and becomes more acidic, they are aquatic ecosystems. Overall, the aluminum dioxide system begins to dissolve, which increases its con-carbon-carbon acid buffer dissolution.tion to keep the pH of water within the field. This is achieved by absorbing dehydrogen ions into the water when they are in excess (that 4.8 Water Movements produce carbon acid and bicarbonates) and through the environment of fresh water and marine production of them when they are scarce (which produces carbonate and bicarbonate ions). In neutral condition (pH-7) the movement of water, flow and most co2 currents is presented as HCO3 (Figu - waves open or breaking water mass in 4.9). In a state of high pH, there is more CO32 than the coast, determines the nature of many environments in low pH, where more CO2 is produced in free condition. Water. Current speed forms the shape and the addition or subtraction of CO2 changes the pH as well as the flow structure. The shape and tilt of the exchanges in pH changes CO2. Flow channels, width, depth and hardness of the bottom, the intensity of rains and the speed of thaws - the pH of natural waters varies from 2 to 12. Water, all of which affects speed. Fast streams emanating from geologically dominant limestone basins will have a higher pH and will easily be loving - those whose speed is 50 cm per second or more, compared to the water of domi basins (see quantitative ecology 24.1: River discharge). A nest of sandstone and granite. Having ions of this velocity, the current will remove all highly alkaline particles, potassium and calcium from the ocean waters - less than 5 mm in diameter and leave behind stony-nica soil leading to seawater slightly alkaline. A large volume of water increases speed; line, in the range of 7.5 to 8.4. Moves stones and materials from the bottom, scrubs pH of the water environment is the element that beds and forms new jars and canals. As ve-can exert a powerful effect on distribution and in locity it reduces both the width, depth and volume of the abundance of organisms. Increased water acidity increases, yl and deorganic substances can affect organisms as directly, by the composition accumulating in the background. Thus, the shape of its effect on physiological processes, as in the flow changes depending on the speed of the current, direct, affecting the concentration of metals weighs- The wind generates waves in large lakes and in Toxic. PH tolerance limits vary depending on the discovery. Resistance to wind surfaces clears plant or animals, but most calm aquatic organisms cause them to produce ripples. They are unable to survive and multiply at a lower pH as it continues to blow, the wind generates more pressure up to 4.5. A factor that contributes greatly to the inability of the more sloping side of ripples, and they aquatic organisms tolerate pH-sized wave conditions starting to grow. As winds high concentrations of aluminum waters become more intense, severed and short decadal waves appear. Aluminium is highly toxic to many speci-of-all sizes that continue to grow as they absorb aquatic life and therefore it generates more energy. When the waves reach a point where the total population is in the aquatic population, the energy that wind supplies is equal to the energy that Aluminium is insoluble when the pH is neutral or basic. lost in the switch, become waves with foam inDevable aluminum represents in high concentrations of ridges. The stronger the rocks, soils, rvers and wind deposits, the higher the waves, the lakes. Under normal pH conditions, concentration-waves that break down on the beach do not form aluminum in the water lakes are very low. There is no water coming from distant seas. Each particle dewwww.FreeLibros.org-Chapter4 Ambient 81 81 ecologia 6ta edicion smith pdf gratis

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