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The study of environmental chemistry plays an important role. The impact of undesirable changes occurring in the environment has a negative impact on humans, plants and animals. Atmospheric pollution occurs in all three states of matter. Note: Under the revised CBSE curriculum, the full chapter has been removed from the curriculum for the 2020-21 academic session. There are various strategies for studying stratosphere pollution and tropospheric pollution. The lowest area of the atmosphere is the troposphere. There are organisms such as humans, animals, plants, etc. One important component of this layer is ozone. The depletion of the ozone layer is gradually beginning to decrease due to industrial carbonated chemical compounds such as bromine or chlorine, as well as various human activities. This is called ozone depletion. Read more: Greenhouse effect Troposphere, which is the lowest layer of the atmosphere, as well as the Earth's surface, is heated due to the existence of carbon dioxide, water vapor, methane, etc. Read more: Greenhouse Effect Several important issues What is the biochemical demand for oxygen? Explain in detail. Identify the smog? The difference between classical smog and photochemical smog? Which gas is more dangerous? Carbon monoxide gas or carbon dioxide gas? Explain why? List the harmful effects of photochemical smog and explain the measures to combat the same. For more information on environmental chemistry, watch the video below: To learn more about this chapter and download the Environmental Chemistry Class 11 Notes PDF Register from BYJU'S. Other important references: Chemistry Notes for Class 11 Chemistry Environmental Chemistry is an important chapter for class 11 students as it deals with environmental topics from a chemical perspective. The chapter also talks about the study of reactions, effects, origins, and mainly the fate of species in the environment. In addition, the themes are huge and students can't remember all the details that can interfere with their performance. Keeping that mind up, our experts have created CBSE revision notes for Class 11 Chemistry Chapter 14 - an environmental chemistry that provides solid and solid knowledge on topics. The information provided is even more relevant and substantive, and these notes will serve as an important reference tool for a clear understanding of the chapter. Students can also learn more effectively and score more in exams. CBSE Class 11 Chemistry Chapter 14 Environmental Chemistry Also Read: Full guide to CBSE Students NCERT Solutions, NCERT Notes, Free Videos, CBSE Documents, TSTs MCQ and more. 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Revisiting notes on exam days is one of the best tips recommended by teachers during exams. Review Notes to Class 11 Chemistry Download as PDF CBSE Class 11 Chemistry Notes Chapter 14 Environmental Chemistry Air, Water and Soil Pollution Chemical Reactions in the Atmosphere Smog, Acid Rain ozone and its Reactions, Green House Effects and Global Warming Strategia to Control Pollution Antirional Chemistry plays an important role in the environment. Chemical species present in the environment are either natural or generated by human activities. Pollution is a consequence of undesirable changes in our environment that have harmful effects on plants, animals and people. Pollutant: - The substance that causes pollution is known as a pollutant. Pollution can be solid, liquid or gas-free substances present in greater concentration than in nature and are produced as a result of human activity or as a result of natural phenomena. For example, substances such as dichlorodiphenylchlor lubricant (DDT), plastic materials, heavy metals, nuclear waste, etc., are difficult to remove once released into the environment. Pollutants can be biodegradable and non-degradable: Biodegradable pollutants: These are pollutants that are rapidly decomposed by natural processes. Example: Discarded vegetables - biodegradable pollutants: These are pollutants that slowly decompose and remain in the environment in an unchanged form for many decades. Pollution has three types: a. Atmospheric pollution. Tropospheric pollution ii. Stratospheric pollution Atmospheric pollution is commonly studied as tropospheric and stratospheric Atmospheric pollution occurs when the concentration of a normal air component or new chemical added or formed in the air is formed to undesirable proportions, causing harm to humans, other animals, animals, Materials. Troposphere: The lowest area of the atmosphere in which humans live together with other organisms is called the troposphere. It extends to a height of 10 km from sea level. The troposphere is a turbulent, dusty zone containing air, lots of water vapor and clouds. Tropospheric pollution: Tropospheric pollution occurs due to the presence of unwanted solid or carbonated particles in the air. Below are the main carbonated and particulate pollutants present in the troposphere: 1. Gaseous air pollutants: These are sulphur, nitrogen and carbon oxides, hydrogen sulfide, hydrocarbons, ozone and other oxidizers. 2. Pollutant particles: This is dust, fog, smoke, smog, etc. Acid rain: Normally rainwater has a pH of 5.6 due to the presence of ions of NK, formed by the reaction of rainwater with carbon dioxide present in the atmosphere. H2O (l) - CO2 (g) → H2CO3 (aq) Source: Burning fossil fuels (which contain sulfur and nitrogen), such as coal and oil in power plants and furnaces, or gasoline and diesel in motor engines, produce sulphur dioxide and nitrogen oxides. SO2 and NO2 after oxidation and water reaction are the main factors of acid rain because contaminated air usually contains particulate matter that catalysis oxidation. Harmful effects: -Harmful to agriculture, trees and plants as it dissolves and washes away the nutrients needed for their growth. Causes respiratory diseases in humans and animals. It affects the life of plants and animals in the aquatic ecosystem when acid rains fall and flow as groundwater to reach rivers, lakes, etc. Corroding water pipes, which leads to the leaching of heavy metals such as iron, lead and copper, into drinking water. Damages buildings and other structures made of stone or metal. The Taj Mahal in India has been hit by acid rain. Gaseous pollutants descend on the ground in the form of acid rains. Green house effect: About 75% of the solar energy reaching the Earth is absorbed by the Earth's surface, which increases its temperature. The rest of the heat radiates back into the atmosphere. Some of the heat is trapped in gases such as carbon dioxide, methane, ozone, chlorofluorocarbon compounds (CFCs) and water vapor in the atmosphere. Thus, they add to the heating of the atmosphere. This leads to global warming. It captures the heat of the sun near the Earth's surface and keeps it warm called the natural greenhouse effect. It maintains temperature and makes the Earth ideal for life. 75% of the solar energy reaching the Earth is absorbed by the Earth's surface, and rest is emitted back into the atmosphere. These gases mentioned above trap heat that lead to global warming. It is important to understand that these gases are also responsible for life on Earth, because they are the amount of solar energy needed to sustain life. Increase Increase gases increase the temperature of the Earth's atmosphere, which, if not checked, can eventually lead to the melting of polar ice caps and, therefore, can lead to flooding of the value mass of the land. Smog: Smoke is a mixture of smoke, dust particles and small droplets of fog. Classic smog Photochemical smog occurs in cool humid climates. This occurs in warm, dry and sunny climates. It is a mixture of smoke, mist and sulphur dioxide. It is also called smog reduction. It is also called oxidizing smog. Stratosphere: Above the troposphere is the stratosphere at a depth of 10 to 50 km above sea level. The ozone layer is an important component of the stratosphere. The presence of ozone in the stratosphere prevents about 99.5 percent of harmful ultraviolet (UV) radiation from reaching the Earth's surface and thus protects humans and other animals from its effects. The formation and fission of ozone Top stratosphere consists of a significant amount of ozone (O3), which protects us from harmful ultraviolet (UV) radiation (255 nm) coming from the Sun. UV radiation breaks down molecular oxygen into free oxygen (O) atoms. These oxygen atoms are combined with molecular oxygen to form ozone. O2 (g) -UV- → O (g) - O (g) O (g) - O2 (g) ≠ O3 (g) Ozone thermodynamically unstable and decomposes into molecular oxygen. Thus, there is a dynamic balance between the production and decomposition of ozone molecules. Many human activities produce chemicals that cause the depletion of the ozone layer in the stratosphere, leading to the formation of an ozone hole. Ozone depletion is the main cause of ozone depletion, and the release of chlorofluorocarbon (CFC) compounds, also known as freons, is considered to be the main cause of ozone depletion. These compounds are used in refrigerators, air conditioners, plastic foam and electronics to clean computer parts, etc. once CFCs are released into the atmosphere, they are mixed with normal atmospheric gases and eventually reach the stratosphere. In the stratosphere, they break down powerful ultraviolet radiation, releasing free radical chlorine. CF2Cl2 (g) → UV → Cl (g) + CF2Cl (g) Cl (g) + O3 (g) → ClO (g) + O2 (g) ClO (g) + O2 (g) → Cl (g) + O2 (g) ClO (g) + O2 (g) → Cl (g) + O2 (g) Thus, chlorine radicals are continuously regenerated and cause the ozone layer to disintegrate. Thus, CFCs transport agents for the continuous production of chlorine radicals into the stratosphere and damage to the ozone layer. The effects of ozone depletion lead to skin aging, cataracts, sunburn, skin cancer, many phytoplankton, and productivity damage. It has also been reported that plant proteins are easily affected by UV radiation, UV radiation, harmful cell mutations. It also increases the evaporation of surface water through the leaf stomata and reduces moisture content in the soil. Increased UV radiation damages paints and fibers, causing them to disappear faster. Water is the elixir of life, but the same water, if contaminated with pathogens, organic waste, toxic heavy metals, pesticides, etc., will turn into poison. WATER POLLUTION Water is essential for life. Water pollution occurs as a result of human activities. Through various pathways, pollution reaches the surface or groundwater. A easily identified source or place of contamination is called a point source. For example, municipal and industrial discharge pipes where pollutants enter the water source. Non-point sources of pollution are those where the source of pollution cannot be easily identified, such as agricultural sources (from farm, animals and farmland), acid rains, storm drains (from streets, parking lots and lawns), etc. Causes water pollution (i) Pathogens: The most serious water pollutants are diseases called pathogens. Pathogens include bacteria and other organisms that enter the water from household wastewater and animal excrement. Human excrement contains bacteria such as Escherichia coli and Streptococcus faecalis, which cause gastrointestinal diseases. (ii) Organic waste: Another major pollutant of water are organic substances such as leaves, grass, garbage, etc. Excessive growth of phytoplankton in water is also a cause of water contamination. These wastes are biodegradable. The main pollutants of water Pollutant Source - Organisms Posal wastewater Organic waste Postal wastewater, animal excrement and waste, decomposing animals and plants, discharge from food plants. Plant Nutrients Chemicals Toxic Heavy Metals Industrial and Chemical Plants Erosion Soil Farming and Strips mining Pesticides Chemicals used to kill insects, fungi and weeds Biochemical Oxygen Demand (BOD): The amount of oxygen needed by bacteria to break the organic substance present in a certain volume of water sample called BOD. Eutrophication: A process in which nutrient-rich reservoirs support a dense plant population that kills animal life, depriving them of oxygen and leading to subsequent loss of biodiversity, known as eutrophication. International Standards for Drinking Water For: For drinking purposes, water must be tested for the concentration of fluoride ions. Its lack of drinking water is harmful to humans and causes diseases such as tooth decay, etc. add to drinking water to bring its concentration to 1 ppm or 1 mg of dm-3. Lead: Drinking water is contaminated with lead when lead pipes are used to transport water. The prescribed concentration of the upper limit of lead in drinking water is about 50 ppb. Lead can damage damage liver, reproductive system, etc. Sulfate: Excessive sulfate in drinking water causes a laxative effect, otherwise at moderate levels it is harmless. Nitrate: The maximum limit of nitrate in drinking water is 50 ppm. Excess nitrates in drinking water can cause diseases such as mezenogloblinemia (blue baby syndrome). Other metals: Maximum concentration of some common metals is recommended in drinking water. Fe: 0.2 ppm Al: 0.05 ppm Mn: 0.2 ppm Cu: 3.0 ppm Ni: 5.0 ppm Cd: 0.005 ppm Therefore, you need to take care to follow international standards to maintain the purity of drinking water. Industrial waste and the excessive use of pesticides contaminate land and reservoirs. Pollution control strategies can be: waste management, i.e. waste reduction and proper disposal, as well as recycling of materials and energy as well. Waste management separates waste as biodegradable and nondegradable waste: biodegradable waste: generated by cotton mills, food plants, paper mills and textile factories. Management: put to landfills and converted into compost. Non - biodegradable waste: generated by thermal power plants that produce fly ash; integrated steel mills that produce blast furnace slag and steel smelter. Management: - Recycling - Toxic waste is usually destroyed by controlled burning environmental chemistry notes for class 11 pdf. environmental chemistry class 11 notes for jee. environmental chemistry class 11 notes for neet

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