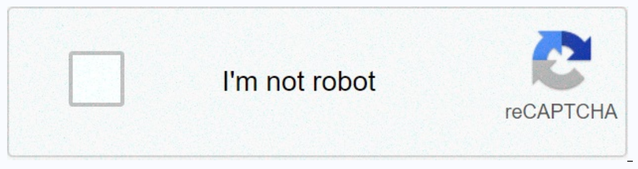

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Molar volume of hydrogen gas lab report

This preview does not show reading a free preview page 2. The laboratory report determines the size of the mole from the gas prologue: This laboratory is designed so that we students can learn how to effectively determine the size of the gas molar. Method: The first step we took to achieve our goal was to put our safety glasses and choose a laboratory station to work in. We received one 400ml beak, one polyethylene pipe, two test tubes with pit rubber plugs, two small pieces of magnesium (mg), one thermometer and a vial of hydrochloric acid (HCl). We took a 400ml beak and filled about 2/3 full of water (H20) which was 18 OC. Then we measured our pieces of mg at 1.5 cm, and decided that their mass was 1.36* 10-2 g. We filled the pipe 2/3 full of HCl and poured it into one of the test tubes. Then, we covered HCl with h2O just enough so that H2O would not be offset when a stopper was introduced. After inserting the stopper, we put a mg strip into a hole, inverted test tube and put it in a 400ml beak. HCl is heavier than H2O, so it is kinder than the tube, at the bottom of the quart, interacting with mg along the way to produce hydrogen gas (H2). Then we measured the size of H2, cleaned our equipment and performed the experiment for the second time. Results: Formula for reducing H2 volume to STP in trial 1: P1 V1 = P2 V2 766.0 mmHg * 14.5ml = 760 mmHg × T1 T2 291 K 273 K X = 13.71 ml equation to find molar size in trial 29 1 1 with conversion to liters: Size H2 in STP 13.71 ml * 1L moles of mg used 5.6* 10-5 mole sized moles in trial 1 = 24.48 liters / mole when interacted with Mg HCl, we noticed a raging reaction. Discussion: The mall size of H2 in our experience is very close to the size of theoretical moths, but I think the deviation lies in the temperature of H2O: in the first experiment it is very high and in a very low second. The purpose of this laboratory is to determine the amount of gas produced in reaction by measuring the change in pressure on the system. Then, we had to identify the mall volum of the gas produced. Products / Reactants name FormulaConcentration / Amount of molar mass used / producing chemical properties properties of chemicals consomes magnesium tape, hydrochloric acid mg, HCl24.31 g/mol (mg)3 m (HCl).02/.08/.1 2 g Mg3/6/7 ml HClMg - BP = 1100 C MP = 651 CHCl- BP = -85.1 C Pvap = 42.6 barMg- hexagonal crystal packed, flexible when hot, strong cut agent. Strong HCl acid, corrosion of metals and tissues mcchlorium products, hydrogen gas 2, H2MgCl2-MP = 118 C BP = 1412 C H2- Density: 0.0899*10-3 g.cm-3 at 20 CBoiling points: - 252.8 CMelting Points: - 259.2 C MgCl2 - Colorless, Crystalline, Soluble in WaterH2- Forms Bonds, Diatomic Atom, Ionization Energy 1:1311 KW-1111 186.59 Trial 3 block of mg (g) 0.020.08 Vial volume (ml) 157157 1577 Maximum pressure (KPa) 142.5179.44 196.59 Initial pressure (k.pa) 10098.89.88 Pressure change (kPa) 42.580.64 96.71 Temperature (K) 298298.3 298.7 Steam pressure from water (kPa) 3.167 3.224 3.302de/conclusion once we had collected all our data, we were able to find the molar size of H2 produced in our reactions. We used the first sober measurement to calculate the number of moles of magnesium tape used. In Reaction 1, we used 8.2304*10^-4 mols. The work is presented below. (0.02 mg/24.4 g/mol) = 8.2304 * 10^-4 mol since there is 1 H2 per 1 mg in this reaction, we can say that there will be 8.2304 * 10^-4 MMol H2 produced. Then we used the Pv=nRT equation to find the Size of H2 produced in reaction 1.Pv = nRT (142.5 kPa) (v)=8.2304 * 10^-1 4 mall) (8.314) (298 K) v = 0.01431 L H2 and then we divided the H2 volume on moles from H2 to find the volume of the mall. (0.01431 L H2) / (8.2304* 10^-4 Mol H2) = 17.3856 L/mol using this molar mass and common gas law, we converted the mall block to STP (101.3 KPa and 273 kK) P1V1/T1 = P 2V2/The government's policy of supporting the government of the united states of the world is to be a key component of the united states' policy of the right to development and to support the development of the united states of the world. We were able to conclude that our results were very similar to the accepted value, which meant that we did something right and therefore we were very happy. Happy.

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