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Thank you so much for your cooperation. The laws of thermodynamics form the basis for all energy-sharing systems. Explore these processes and identify the fundamental physical quantities that define the world around us. Learn about ideal gases, theoretical gases that adhere to the ideal gas law, which theoretically is an impeccable mathematical link between pressure, volume and temperature. Explore non-idealized gases whose molecules are bound by attractive forces that make them behave more realistically. Dig into the application of the zero law of thermodynamics, which states that if two thermodynamic systems are in balance with the third system, they are in balance with each other. Apply your knowledge of the first law of thermodynamics, applying the law of energy conservation to solve for heat, entropy and other variables. The second law of thermodynamics states that entropy in an isolated system increases over time. Use this knowledge to determine the characteristics in systems and calculate changes in entropy and efficiency. The third law of thermodynamics, developed by Walter Nernst, states that as the system approaches absolute zero, the entropy changes the system undergoing isothermal process reaches zero as well. What drives the physical or chemical process? Deep into the concepts of free energy and chemical potential to determine the conditions of spontaneity and the characteristics of the state of equilibrium. How do changes in free energy make the process spontaneous or not spontaneous? Compare and compare the free energies of Gibbs and Helmholtz and how they are used to produce many thermodynamic mathematical relationships. The myeregime methods of partial differentiation and integration to obtain a link between thermodynamic properties. Get a higher appreciation for the application of thermodynamics laws because you use Maxwell's relationship to evaluate properties for real and ideal systems. Fugacity is a measure of how likely gas is to expand Escape; apply this property to simulate the real behavior of the gas. Laws of Laws thermodynamics eventually indicate a state of equilibrium. Understand and know the quantitative treatment of balance and factors that influence balance. Explore equilibrium involving gashistory, axial and condensed phases, as well as equilibrium in single- and multi-component systems. Browse the chemical equilibrium, where reactionions and products are present in constant ratios. Analyze how the chemical equilibrium occurs and stabilizes in solutions in a compressed phase. The balance of constants, unlike their names, can vary depending on systemic conditions, such as temperature changes. Chemical equilibria is widespread in living systems - learn how they are applied to biochemical reactions associated with amino acid pathways. Focus on phase transitions, Clapeyron equations, and gas phase effects in the simplest systems: a single-component system. Most chemical systems have multiple components. Immersion in the Gibbs phase rule, colligative properties and various types of equilibrium multicomponent systems: liquid/liquid, liquid/gas, liquid/solid and solid/solid solutions. The transition from the classical world to a world in which the quanta dictates the interaction of matter and energy. Discover the origins of the quantitative nature of electronic transitions and understand the behavior of atoms controlled by the laws of quantum mechanics. Review the fundamental laws of Isaac Newton's movement that encapsulate the movement of the bodies of matter. Analyze the unexplained phenomena that have confused classical scientists: atomic line spectra, atomic structure, light and photovoltaic effect. Use the Rydberg equation to calculate wavelengths and wave intentions of light. Learn about the early discoveries of Heinrich Hertz and Philippe Eduard Anton von Lenard regarding the photovoltaic effect. Scientists have been studying and discussing light for centuries. Analyze experiments with historical light, including Thomas Young's double-cut experiment. Apply the laws of Max Planck and Albert Einstein to determine the energy of light. Assess how Niels Bohr repurposed Rydberg's equation to develop atomic atomic theory and quantify measurable quantities. Louis de Brogley suggested that particles such as electrons may have wave properties. Using the de Brogley equation, predict the relationship between wavelength and momentum. The quantum mechanics forced scientists at the turn of the century to consider a new version of reality. Explore the fundamental principles of quantum mechanics, from

the mysterious nature of wave particle duality to probabilistic wave function and how it controls the behavior of our microscopic world. You will find the truth of quantum mechanics to be stranger than science fiction. Explain how wave behavior can be expressed mathematically to identify expressions that are wave-we're going to have to do. wave-we're going to have to do. values observed, and use different operators to study eigenvalues and eigenfunctions. Use the Heisenberg Uncertainty Principle to measure uncertainty in different contexts. Based on the Heisenberg Uncertainty Principle, use the Born Interpretation of Wave Function to calculate the probability of where the electrons are. Normalization of wave functions in accordance with Bourne's interpretation of wave function. Erwin Schroedinger's equation is used to determine the significant observed in quantum mechanics: energy. Calculate the averages (or expectations) of the different observed, and predict where and when tunneling may occur in the system. Apply the concept of degeneration, where different linearly independent wave functions have the same energies with different sets of quantum numbers. If two wave functions are multiplied together and their integral is zero, they are considered orthogonal. Demonstrate through evidence how wave functions can be orthogonal to each other. Explore various model systems in quantum mechanics, including classic harmonic oscillator, two- and three-dimensional rotations, and wave functions of a hydrogen atom. Assess how symmetry is seen in different wave functions. What are the basics of chemistry? Chemistry explains everything that is based on atoms and molecules. Atoms and molecules are explained by quantum mechanics. Explain the importance of the innate angular impulse, which all electrons are known as spin. Apply the basic principles of quantum mechanics to develop models for atomic helium, as well as some simple diatom molecules. According to the Pauli principle, each electron in this system has a unique set of four quantum numbers. Use this principle to determine spin orbits for electrons. Assign electrons to rotate orbits in more complex, larger atoms. The perturbation theory states that the real system can be approximated as a known system and that any violation can be calculated separately from the system and later added. The theory of variations states that for any test wave dysfunction for the system the energy is greater or equal to the energy of the terrestrial state. Dig out the approach of Born-Oppenheimer, the most important base for molecular quantum mechanics. Using the theory of linear variations, use linear combinations of occupied atomic orbital stations to synthesize molecular orbital stations. Apply the basic principles of quantum mechanics to the development of models of simple diatom molecules. Start studying symmetry with group theory: symmetry and group operations, character tables, great orthogonality theorem, symmetry-adapted linear combinations, and valence theory. What are the practical applications of quantum mechanics? Spectroscopy. Using quantum-mechanical models we develop the basic principles of rotation (microwave) and vibration (infrared) spectroscopy. Spectroscopy is instrument of this course of physical chemistry. Determine the number of vibrational degrees of freedom for different molecules. Explore the kinetic area, or how and how quickly the reaction occurs. Master the derivatives and usefulness of speed equations for reactions on different orders. Learn about different reaction mechanisms, including enzyme catalysis, parallel, consistent, chain and oscillating reactions, and theories governing reaction rate. See the speed at which chemical reactions occur with speed and betting laws. Explore how catalysine enzymes chemical reactions, how reaction rates are measured experimentally, and how different reaction conditions affect speed. Learn how catalysts accelerate the speed of reaction, reducing the energy of activation. When the reaction approaches equilibrium, the concentration of reactionary means and products begin to equalize the score. Compare and compare parallel and consistent reactions. Parallel reactions result in when more than one product is possible, and consistent reactions use the products as a reaction in subsequent reactions. Learn how temperature affects the rate of chemical reactions and how energy activation stems from temperature changes and enthalpy. The stable state of approximation takes into account the fact that intermediates often occur in reactions, but are neither a reaction nor a product. Assess how the approximation is used to get speed laws. Compare and contrast chain and fluctuating reactions and how they are on the chart in relation to temperature and pressure. The theory of the transitional state (also known as the theory of activated complex) simulates bimolecular elementary steps in the reaction. Look at the basic definitions and properties of gases, as well as the basic laws on gas and how collisions occur. Occur.

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