## Empirical and molecular formula worksheet with answers

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Write the empirical formula for the following compounds. 1) C6H6 CH 2) C8H18 C4H9 3) WO2 WO2 4) C2H6O2 CH3O 5) X39Y13 X2Y 6) A compound with an empirical formula of C2OH4 and a molar mass of 88 grams per mole. What is the molecular formula of this compound? First, find the empirical formula mass of the compound. Carbon: 2 (12.0 g/mole) Oxygen: 1(16.0 g/mole) Hydrogen: 4(1.0 g/mole) Hydrogen: 4 as the empirical formula. 88/44 = 2 Molecular formula is 2 times the empirical formula 2 (C2OH4) = C4O2H8 7) A compound with an empirical formula of this compound? Use the same strategy as problem #6 4 (12 g/mol) + 4(1 g/mol) + 1 (16 g/mol) = 68 g/mol) = 68 g/mol mol 136/68 = 12 2 (C4H4O) = C8H8O2 8) A connection to an empirical formula of CFBrO and a molar mass of 254.7 grams per mole. What is the molecular formula of this compound? Use the same strategy at problems 6 & amp; 7. 1(12 g/mol) + 1(19 g with an empirical formula of C2H8N and a molar mass of 46 grams per mole. What is the molecular formula and the empirical formula the same. 10) A well-known reagent in analytical chemistry, dimethylglyoxime, has the empirical formula C2H4NO. If its molecular mass is 116.1 g/mole, what is the molecular formula of the compound? C:2(12 g/mol) H:4(1 g / mol) N:1(14 g/mol) O: 1(16 g/mol) = 24 + 4 + 14 + 16 = 16 = 58 g/mol 116.1/58 = 2 (C2H4NO) = C4H8N2O2 12) Nitrogen and oxygen form an extensive range of oxides with the general formula NxOy. One of them is a blue solid that comes apart, reversible, in the gas phase. It contains 36.84% N. What is the empirical formula of this oxide? Predecion that you have 100 grams of this substance. That way, you can convert all percentages directly to grams. Divide the mass into grams by the atomic mass to determine the number of moles of each atom. 36.84% N = 36.84% N empirical formula of N2O3 . 13. A sample of indium chlorid weighing 0.5000 g is found to contain 0.2404 g chlorine. What is the empirical formula of the indium should have a mass of 0.5000 g - 0.2404 g = 0.2596 g Cl: 0.2404 g / (35.5 g/mol) = 6.77 x 10-3 mole Ir: 0.2596 g / (192.2 g/mol) = 1.35 x 10-3 mole So we have Ir1.35E-3Cl6.77E-3 to get integers, divided by the smallest of these numbers, 1.35 x 10-3. This leaves us with IrCl5. (tip: when you split, you can only do 6.77 / 1.35 and forget the powers of 10. Since they are both 10-3, they cancel.) 14. An unknown compound was found to have a percent composition as follows: 47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its molecular formula? K:47 g / (39 g/mol) = 1.21 mole C:14.5 g / (12 g/mol) = 1.21 mol O: 38.5 g / (16 g/mol) = 2.41 mole Split each number by the smallest to get KCO2 15. Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of rub alcohol? The oxygen should be 100% – 60% - 13.4% = 26.6% Imagine we have 100 g rub alcohol. Now we can convert these percentages to grams. Then we can split by getting their atomic mass mole. C:60 g / (12.0 g / mol) = 5 mole H: g / (1.0 g/mol) = 13.4 mole O: 26.6 g / (16.0 g/mol) = 1.6625 Split each number according to the smallest number, 1.6625. It gives us C3H8O. C3H8O.

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