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Science is a topic that explains how the world around us is made up. Chemical reactions are used to explain the different processes that occur around us. From rust to decomposition, chemical reactions provide a deeper understanding of how molecular interaction and change occur. In other words, Chapter 1 of Class 10 CBSE Science explains how a substance changes shape. Learn more about chemical reactions and equations by researching CBSE Notes for Class 10 Science Chapter 1. These CBSE notes are comprehensive and detailed, but brief enough to look in to prepare for the exam. Chemical change - one or more new substances are formed with new physical and chemical properties. Here, when copper sulfate reacts with iron, two new substances are formed, i.e. sulfate and copper. Physical change - color or condition changes occur, but a new substance is not formed. Example: Water changes to steam when boiled, but a new substance is not formed (Even if steam and water look different when they are made to react with a piece of Na, they react in the same way and give exactly the same products). This only includes a change in condition (liquid for steam). Observations that help determine the chemical reaction Chemical reaction can be determined by using any of the following observations: a) Evolution of gas b) Temperature change in) Formation of precipitation (d) Changing the state of chemical reaction Chemical reactions chemical changes in which the reaction is converted into products by creating or severing bonds (or both) between different atoms. Types of chemical reactions Taking into account different factors, chemical reactions are grouped into several categories. Few examples: Combination - Decomposition - One Move - Double Movement - Redox - Endothermic - Exothermic - Precipitation - Neutralization of Chemical Reactions and Equations I Word Equation The Word Equation is a chemical reaction expressed in words, not chemical formulas. This helps determine the reaction and products in the chemical reaction. For example, sodium and chlorine → sodium chloride the aforementioned equation means: Sodium reacts with chlorine to form sodium chloride. The symbols of the elements and their valence Symbol is the chemical code of the element. Each element has one or two letters of atomic symbol, which is an abbreviated form of its name. Valency is combining the capacity of an item. It can be seen as the number of electrons lost, lost or divided by an atom when it is combined with another atom to form a molecule. Writing chemical equations Representation of chemical reaction in terms of symbols and chemical formulas of reactionary and products is known as the chemical equation. For solids, the symbol is For liquids, it's (l). For gases, it's (g). For The For solution, it's (aq). For the gas produced in reaction, it is represented (C). For precipitation, formed in reaction, it is represented (C). Balancing the chemical reaction Saving mass in accordance with the law of mass preservation, no atoms can be created or destroyed in a chemical reaction, so the number of atoms for each element in the direction of reactionary should balance the number of atoms that are present in the side of the products. In other words, the total mass of products formed as a result of a chemical reaction is equal to the total mass of the reaction that participated in the chemical reaction. A balanced chemical equation is a chemical equation in which the number of atoms of each element in the reactionary side is equal to what of the side's products is called a balanced chemical equation. Steps to balance chemical equations Hit and trial method: When balancing the equation, change the ratios (numbers before the compound or molecule), so that the number of atoms of each element is the same on each side of the chemical equation. Short-term method of balancing the chemical equation Example: aCaCO₃ and bH₃PO₄ → cCa₃(PO₄)₂ and dH₂CO₃ Setting a series of simultaneous equations, one for each element. Ca: a'3c C: a'd O: 3a'4b'8c'3d H: 3b'2d P: b'2c Let's install c'1 then a'3 and d a 3 b 2c 2 So a'3; b-2; c-1; d'3 Balanced Equation 3CaCO₃ and 2H₃PO₄ → Ca₃(PO₄)₂ - 3H₂CO₃ Chemical Reactions and Equation II Types of Chemical Reactions Taking into account different factors, chemical reactions are grouped into several categories. Few examples: Combination of Decomposition Double Move Double Move Redox Endothermic Exothermic Precipitation Reaction combination of neutralization In the combined reaction, two elements or one element and one connection or two connections are combined to give one separate product. The decomposition reaction One responsive decomposes when applying heat or light or electricity to give two or more foods. Types of decomposition reactions: a. Decomposition reactions requiring heat are thermolytic decomposition or thermolysis. Heat decomposition HgO b. Decomposition reactions requiring light - photolytic decomposition or photolise. Photolytic decomposition H₂O₂ c. Decomposition reactions requiring electricity - electrolytic decomposition or electrolysis. Electrolytic H₂O Decomposition For more information on decomposition reactions, watch the video below: The reaction of moving the More Reactive Element displaces a less reactive element from its compound or solution. The dual reaction of moving the exchange of ions between reactions occurs to give new For example, the reaction of precipitation is an insoluble compound called sediment form, when two soluble salts are combined. Redox reactions of oxidation and contraction occur simultaneously. Oxidation: The substance loses electrons or receives oxygen or loses hydrogen. Reduction: The substance acquires electrons or loses oxygen or receives hydrogen. An oxidizing agent is a substance that oxidizes another substance and asserts itself. Agent reduction is a substance that reduces another substance and oxidizes. Endothermic and exothermic reaction of exothermic reaction - heat during reaction. Most combined reactions are exothermic. Al-Fe₂O₃ → Al₂O₃ - Fe - Heat CH₄ , 2O₂ → CO₂ - 2H₂O - Thermal Endothermic - Heat is required to perform the reaction. 6CO₂ - 6H₂O - Solar → C₆H₁₂O₆ - 6O₂ Glucose Most decomposition reactions are endothermic. Corrosion Gradual deterioration of the material is usually a metal, resulting in the action of moisture, air or chemicals in the environment. Rust: 4Fe (s) 3O₂ (from the air) - xH₂O (moisture) → 2Fe₂O₃.xH₂O (rust) Copper corrosion: Cu (s) - H₂O (moisture) - CO₂ (out of air) → → CuCO₃.Cu (OH)₂ (green) Silver Corrosion: Ag (s) H₂S (out of air) → Ag₂S (black) H₂ (g) Rancidity This refers to the oxidation of fats and oils in food that is stored for a long time. It gives a bad smell and a bad taste for food. Rancide food causes stomach infection when consumed. Prevention: (i) The use of airtight packaging containers (ii) with nitrogen (iii) Cold (iv) The addition of antioxidants or Class 10 preservatives is a critical moment for students on their educational journey. In addition, the scores scored on board the exam, Class 10 will also help students choose the desired flow for their higher education. Therefore, it is important that every student effectively learn and score good marks in the CBSE Grade 10 exam board. In order to help students develop the right teaching methodology and succeed in their exams, we provide chapter wise CBSE class 10 Scientific Notes. Students must go through these notes and address them carefully while studying. Class 10 NCERT scientific notes include detailed explanations, as well as illustrative examples, questions resolved, and sample questions for students from NCERT books. CBSE notes would seem to help them get a good grip and a clear understanding of all the important concepts. The style and format used is very simple and in a clear manner, making it easy for students to save as well as memorizing each concept easily for a longer time. The chapter-wise CBSE Class 10 Science Notes Class 10 Science basically includes some important chapters from the NCERT book, both: Chemical Reactions and Equations, Electricity, Magnetic Effects of Electric Current, Acid Bases and Salt, and coordination, reflection and devastation, the world. In the table below, we provided CBSE Class 10 Scientific Notes on all of these topics. Click on the links below to access chapter wise notes of the Class 10 science subject. The benefits of learning from CBSE Class 10 Science Notes Students get a brief overview of all concepts in science notes. Learning through notes helps them during the revision. This saves students time while preparing for the exam. Students will remember all the important topics from each chapter, referring to these notes. These notes will help the student better understand the chapters. In addition, these notes will work as the best revision resource during board exams. In addition to the CBSE Class 10 Scientific Notes, we have also provided important questions and document samples to help students competently prepare for the board exams in the table below. Important questions for Class 10 Science CBSE Class 10 Examples Of Documents Stay tuned to us for more information on CBSE and NCERT. Download BYJU'S App and subscribe to the YouTube channel to access interactive math and science videos. Don't forget how our Facebook page to update new content on our site. We also share useful articles on our Facebook page to help you with your board exam. The questions and funny facts related to class 10 Chemistry Notes will also be shared on our facebook page, so you can ace your chemistry exam. If you find any error or any problem with the notes, please send us an email protected we are working to provide the best resources for your research, your suggestions in this regard will also be highly appreciated. 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