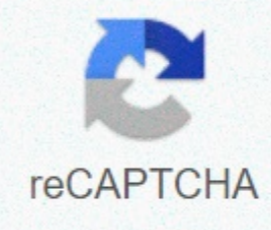




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Department of Chemistry and Physics Worksheet of Arkansas State University Stoichiometry (using solution) 1. Given the following reaction: (hint: balance the equation first) $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$ If 43.2 mL 0.236 M NaOH reacts with 36.7 mL H_2SO_4 , what is the concentration of the H_2SO_4 solution? answer 2. Given the following equation: $\text{NaOH} + \text{HCl} \rightarrow \text{H}_2\text{O} + \text{NaCl}$ If a solution of HCl 36.7 mL is required to react with 43.2 mL NaOH 0.236 M, what is the concentration of the HCl solution? answer 3. Given the following equation: $\text{Al}(\text{OH})_3 + \text{HCl} \rightarrow \text{AlCl}_3 + \text{H}_2\text{O}$ How much mL 1.2 M HCl is required to dissolve 5.8 g $\text{Al}(\text{OH})_3$? answer 4. Given the following equation: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$ Calculating molarity solution H_2SO_4 if needed 40.0 mL H_2SO_4 to neutralize 0.364 g Na_2CO_3 . answer The following chemical reaction balance: a. $2 \text{CO} + \text{O}_2 \rightarrow 2 \text{CO}_2$ b. $2 \text{KNO}_3 \rightarrow 2 \text{KNO}_2 + \text{O}_2$ c. $2 \text{O}_3 \rightarrow 3 \text{O}_2$ d. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + 2 \text{H}_2\text{O}$ e. $4 \text{CH}_3\text{NH}_2 + 9 \text{O}_2 \rightarrow 4 \text{CO}_2 + 10 \text{H}_2\text{O} + 2 \text{N}_2$ f. $\text{Cr}(\text{OH})_3 + 3 \text{HClO}_4 \rightarrow \text{Cr}(\text{ClO}_4)_3 + 3 \text{H}_2\text{O}$ Write a balanced chemical equation of each reaction: a. Calcium carbide (CaC_2) reacts with water to form calcium hydroxide ($\text{Ca}(\text{OH})_2$) and acetylene gas (C_2H_2). $\text{CaC}_2 + 2 \text{H}_2\text{O} \rightarrow \text{C}_2\text{H}_2 + \text{Ca}(\text{OH})_2$ b. When potassium chlorate (KClO_3) is heated, it decays to form KCl and oxygen gas (O_2). $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ c. C_6H_6 burning in the air. $2 \text{C}_6\text{H}_6 + 15 \text{O}_2 \rightarrow 12 \text{CO}_2 + 6 \text{H}_2\text{O}$ d. $\text{C}_5\text{H}_{12}\text{O}$ burning in the air. $2 \text{C}_5\text{H}_{12}\text{O} + 15 \text{O}_2 \rightarrow 10 \text{CO}_2 + 12 \text{H}_2\text{O}$ Given the following reaction: $\text{Na}_2\text{S}_2\text{O}_3 + \text{AgBr} \rightarrow \text{NaBr} + \text{Na}_3[\text{Ag}(\text{S}_2\text{O}_3)_2]$ a. How many $\text{Na}_2\text{S}_2\text{O}_3$ flies does it take to react completely with 42.7 g AgBr? 0.455 mol $\text{Na}_2\text{S}_2\text{O}_3$ b. What is the mass of NaBr to be produced from 42.7 g AgBr? 23.4 g NaBr From reaction: $\text{B}_2\text{H}_6 + \text{O}_2 \rightarrow \text{HBO}_2 + \text{H}_2\text{O}$ a. What O_2 mass would it take to burn 36.1 g B_2H_6 ? 125 g O_2 b. How many water flies are produced from 19.2 g B_2H_6 ? 1.39 mol H_2O Calculate the mass (in kg) of water produced from burning 1.0 gallons (3.8 L) of gasoline (C_8H_{18}). The density of gasoline is 0.79 g / mL. 4.3 kg H_2O One mole of aspartame ($\text{C}_{14}\text{H}_{18}\text{N}_2\text{O}_5$) reacts with two water flies to produce one mole of aspartic acid ($\text{C}_4\text{H}_7\text{NO}_4$), one mole of methanol (CH_3OH) and one phenalanine fly. A. What is the molecular formula of fenalanin? $\text{C}_9\text{H}_{11}\text{NO}_2$ b. What fenalanine mass is produced from 378 g of aspartame? 212 g of phenylalanine KO_2 is used in the respiratory system closed. It removes carbon dioxide and water from exhaled air. The reaction to water removal is: $\text{KO}_2 + \text{H}_2\text{O} \rightarrow \text{O}_2 + \text{KOH}$. The CO_2 KOH produced is used to remove carbon dioxide with the following reaction: $\text{KOH} + \text{CO}_2 \rightarrow \text{KHCO}_3$. A. What mass does KO_2 generate g of O_2 ? 696 g KO_2 b. What CO_2 mass can be removed by 123 g of KO_2 ? 76.1 g CO_2 Department of Chemistry and Physics Arkansas State University Stoichiometry (using solution) 1. Given the following reaction: (hint: balance the equation first) $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$ If 43.2 mL 0.236 M NaOH reacts with 36.7 mL H_2SO_4 , what is the concentration of the H_2SO_4 solution? answer 2. Given the following equation: $\text{NaOH} + \text{HCl} \rightarrow \text{H}_2\text{O} + \text{NaCl}$ If a solution of HCl 36.7 mL is required to react with 43.2 mL NaOH 0.236 M, what is the concentration of the HCl solution? answer 3. 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The reaction to water removal is: $\text{KO}_2 + \text{H}_2\text{O} \rightarrow \text{O}_2 + \text{KOH}$. The produced KOH is used to remove carbon with the following reaction: $\text{KOH} + \text{CO}_2 \rightarrow \text{KHCO}_3$. A. What KO_2 mass produces 235 g O_2 ? 696 g KO_2 b. What CO_2 mass can be removed by 123 g of KO_2 ? 76.1 g CO_2 Some of the worksheets below are Stoichiometry Worksheets with Answer Keys, stoichiometry definitions with lots of examples and interesting exercises involving step by step solutions with some colorful illustrations and diagrams. Basic Instructions After you find the worksheet, you can click the pop-out icon or download button to print or download the worksheet that you want. Please note that you can also find a download button under each document. Stoichiometry Worksheet and Key : Questions such as how many grams of O_2 are needed to react with 125 grams of Fe?, ... 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