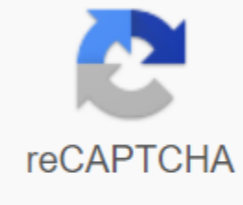




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In stereo-isomeria, or spatial isomeria, there is an optical isomeria that occurs when connections are optically active, i.e. divert the plane of polarized light. Isomers, which show optical activity, are called enantiomers, they have three main characteristics: their molecule is asymmetrical. This means that if we split it in half, the resulting part will not be the same; Enantiomers are a mirror image of each other; They don't overlap, so if we put one molecule above another, they won't be the same; the location of their ligand atoms will be different. These enantiomers, whose molecules are asymmetrical and do not rehash, are called chiral molecules, because the word chiral comes from the Greek cheiral, meaning hand because the hands are asymmetrical and not retooled. In addition, if we put our left hand in front of the mirror, your image will be exactly the same as the right hand, and vice versa. That is, hands are equal to enantiomers, in the sense that they are mirrored by each other. Don't stop now... There's more after the ;) Enantiomers differ from diastereoisomers in that the latter are not a mirror image of each other. Both enantiomers and diastereoisomers are stereoisomers or spatial isomers. However, enantiomers are considered to be the most important, especially with regard to biochemical phenomena. The activity that each enantiomer performs in the body is different from one to the other. For example, the hormone adrenaline has two enantiomers, dextrogyal and levogyro. Dextrogin is often less active as a hormone than levogyro. The effect of the adrenaline levogir is important because it acts as a powerful narrower and hypertensive vessel, which has a pronounced effect on metabolism, causes an increase in heart rate and blood pressure. According to Jennifer Vogania, in Chemistry, the count of the number of diastereoisomers, that is, isomers, which are not mirror images of each other, means the determination of the number of pairs of molecules of optical isomer, which polarize and distract light at different angles. It is noteworthy that when calculating the number of diastereoisomers we do not determine the number of racial mixtures, as this type of mixture is formed only by enantiomorphous vapors (optically active isomers, that is, those molecules that distract light with the same meaning of angulation). Counting the number of diastereoisomers is possible only in substances whose molecules have two or more different chiral carbon, or in open and cyclical chains that have cisterns of geometric isomeria (in these cases we will always have a pair of diastereoisomers). To whom this, we will use 3-methyl-pent-4-en-2-ol, remembering that in this structure, the chain is spotted from right to left, because it is the nearest end of hydroxyl: Structural formula 3-methyl-pent-4-en-en-ol In this compound, there are two chiral carbons that are carbons 2 and 3 structures. Linking each of these carbons: → carbon 2: Demarcation of chiral carbon ligands number 2 and methyl (a); Hydrogen (b); Don't stop now... There is more after the advertising ;) Hydroxyl (c); Sec-no-1-en (h). → 3: Demarcation of carbon queer ligands No. 3 and hydrogen (a); Hydroxiethyl (b); Methyl (c); Vinyl (d). According to the Van't Hoff rule, this compound has four optically active isomers, since: IOA No 2n IOA No 22 IOA No. 4 Therefore, there are four enantiomer molecules in this structure, two dextrogirs (angulation with a positive signal) and two levogiras (an angle with a negative sign). These molecules distract light with the following angularors: - Dextrogir 1: - Levogyro 1: - α - Dextrogin 2: - β Levogyro retell 2: - β From this knowledge, we can count diastereoisomers of the compound 3-2-metyltan-2-ol. Diastereoisomers correspond to each combination of dextrogirs and levogir stumofers with different angles, for example: - Dextrogin 1 (l α) with Levogyro 2 (- β); Dextrogi 1 (l α) with Dextrogiro 2 (β); Levogyro 1 (- α) with Levogyro 2 (- β); Dextrogi 2 (l β) with Levogyro 1 (- α). Thus, the substance 3-methyl-pent-4-en-2-ol represents four optically active isomers, which led to four pairs of diastereur. Therefore, to facilitate the counting of diastereoisomers, we need to know the number of optically active isomers. IOA - The number of pairs of diastereoisomers Published: Diogo Lopes Dias Isomeria occurs when there are two or more compounds that have the same molecular formula, that is, they are formed by the same atoms and in the same amount, but differ in their structure. One type of isomeria is stereoisomeria, also known as spatial isomeria, in which isomers are differentiated only by the spatial arrangement that their atoms acquire in space, i.e. by the configuration of molecules. All stereoisomers belong to the same organic function, have the same type of chain and the same position as insatiable, heteroate and branching (if any). Therefore, it is necessary to analyze its spatial configuration. But there are two types of stereoisomers that are diastereoisomers and enantiomers. See below how to distinguish them and what spatial isomeria they perform: Diastereoisomers: These isomers are not a mirror image of each other, and they perform a deviation of light Open and cyclical chain diastereoisomers exist at different angles. Open chains should have a double bond between the two carbons, and the bonds of each of these carbons should be different, but equal to the other carbon link. See the example below in which we have a 1,2-dichlorone connection: H H Cl H/C? C? C/Cl Cl Cl cis-1,2-dichloroethene trans-1,2-dichloroeteten Not stopping now... There's more after the ;) Note that they are not a mirror image of each other, meaning that when they are located side by side, they do not collide with the mirror. The difference is that equal ligands are on one side of the plane in one (cis diastereoisomer) and on opposite sides of the plane in the second isomer (trans diastereoisomer). Closed circuit diastereoisomers do not have a double bond, but at least two cycle carbon atoms must have different ligands and equal to another carbon atom, as in the example below: Enantiomers: These isomers are a mirror image of each other but do not intersect. In addition, they perform optical isomeria, which when the connection distracts the plane of polarized light. One way to determine whether the compounds are enantiomers is to check that the molecule is asymmetrical because it has at least one chiral carbon, that is, that has all the different ligands. When a molecule obeys this rule, being asymmetrical, it has enantiomers, which, when put side by side, give the impression that they are in front of the mirror, because they are exactly the image of each other. See the example below: Published: Jennifer Rocha Vargas Foganya This page cites reliable sources, but which do not cover all content. Help insert links. Unverified content can be removed.- Find sources: Google (news, books and academic) (September 2014) Diastereoisomers are stereoisomers who are not enantiomers, that is, they are stereoisomers where one of them is not an image in the mirror of the other. Diastereoisomers have different physical properties, and have different chemical properties, both reacting with chiral and achiral compounds. Typically, n binary stereogenic elements (R and S, E, M and P, etc.) lead to 2^n stereoisomers (and therefore up to $2^n - 1$ racial mixtures). An example would be menthol, has 3 chiral centers, and therefore 8 different stereoisomers. Of the 28 possible pairs, only 4 enantiomers, the remaining 24 pairs of diastereoisomers: Menthol. However, the maximum number of stereoisomers 2^n , decreases when the structure has symmetry elements (the most common of which is the reflection plane or the center of the inversion). An example of this is meso-isomers. Such substances are characterized by the presence of two centers of chirality (asymmetric carbon) equal. Meso-isomers: - They have two equal chiral carbon: Suppose that each asymmetrical carbon in this meso develops a plane of polarized light in a degree. Now, according to the convention, for is a deviation to the right (dextrogiorogorium) and -a represents a deviation to the left (levogyro). In this case, there will be the following combinations: 1st (a-a-a) / 2nd (a-a) / 3rd (For) and racial mix, i.e. 4th (a-a) mixed with (For) They are optically active: 1st and 3rd Optically inactive: 2nd and 4th links - IUPAC, Compendium chemical terminology, 2nd ed. (Golden Book). Compiled by A.J. McNaught and A. Wilkinson. Blackwell Scientific Publications, Oxford (1997). Online version: diastereoisomerism (2006-), created by M. Nick, J.Girat, B. Kosata; updates by A. Jenkins. ISBN 0-9678550-9-8. Colle bibliography, A.; J. Crassus; Dutasta, J.-P.; Guy, L. Chirales Molecules. 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